AP Chemistry Final Study Guide

Unit 1: Atomic Structure and Properties (7%-9%)

- At STP, one mole of gas always occupies 22.4 L
- Mass Spectroscopy:
 - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
 - The most abundant isotope will have the greatest effect
- Coulomb's Law:
 - The greater the charge of the nucleus, the more energy is required to ionize an electron
 - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- · Electromagnetic Radiation:
 - When atoms absorb energy, electrons jump to higher energy levels
 - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- Kinetic Energy of Ionized Electrons:
 - The faster an ejected electron is going, the more kinetic energy it has
 - Electrons originally further away from the nucleus require less energy to eject, so
 it will have more kinetic energy
- · Electron Configuration:
 - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
 - Nickel: 1s²2s²2p⁶3s²3p⁶3d⁸4s²(notice the d-block)
 - Nickel (with shorthand): [Ar]4s²3d⁸
- Law of Definite Proportions:
 - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- Law of Multiple Proportions:
 - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- Aufbau Principle:
 - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

Chem Study Guide For Spring Benchmark

D Keegan

Chem Study Guide For Spring Benchmark:

Environmental Health Literacy Symma Finn, Liam R. O'Fallon, 2018-09-12 This book explores various and distinct aspects of environmental health literacy EHL from the perspective of investigators working in this emerging field and their community partners in research Chapters aim to distinguish EHL from health literacy and environmental health education in order to classify it as a unique field with its own purposes and outcomes Contributions in this book represent the key aspects of communication dissemination and implementation and social scientific research related to environmental health sciences and the range of expertise and interest in EHL Readers will learn about the conceptual framework and underlying philosophical tenets of EHL and its relation to health literacy and communications research Special attention is given to topics like dissemination and implementation of culturally relevant environmental risk messaging and promotion of EHL through visual technologies Authoritative entries by experts also focus on important approaches to advancing EHL through community engaged research and by engaging teachers and students at an early age through developing innovative STEM curriculum The significance of theater is highlighted by describing the use of an interactive theater experience as an approach that enables community residents to express themselves in non verbal ways Water quality in the Great and Little Miami River Basins, Ohio and Indiana, 1999-2001, Chemists' Guide to Effective Teaching Norbert J. Pienta, Melanie M. Cooper, Thomas J. Greenbowe, 2005 For courses in Methods of Teaching Chemistry Useful for new professors chemical educators or students learning to teach chemistry Intended for anyone who teaches chemistry or is learning to teach it this book examines applications of learning theories presenting actual techniques and practices that respected professors have used to implement and achieve their goals Each chapter is written by a chemist who has expertise in the area and who has experience in applying those ideas in their classrooms This book is a part of the Prentice Hall Series in Educational Innovation for Chemistry Performance Evaluation and Benchmarking with Realistic Applications Rudolf Eigenmann, 2001 The book discusses rationales for creating and updating benchmarks the use of benchmarks in academic research benchmarking methodologies the relation of SPEC benchmarks to other benchmarking activities shortcomings of current benchmarks and the need for further benchmarking efforts Performance evaluation and benchmarking are of concern to all computer related disciplines A benchmark is a standard program or set of programs that can be run on different computers to give an accurate measure of their performance This book covers a variety of aspects of computer performance evaluation with a focus on Standard Performance Evaluation Corporation SPEC benchmarks SPEC is a nonprofit organization whose members represent industry academia and other organizations The book discusses rationales for creating and updating benchmarks the use of benchmarks in academic research benchmarking methodologies the relation of SPEC benchmarks to other benchmarking activities shortcomings of current benchmarks and the need for further benchmarking efforts Contributors Brian Armstrong Frederica Darema Edward S Davidson Sylvia Dieckmann Jozo J Dujmovic Rudolf

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