

Effect of Concentration on the Reaction Rate: Finding the Rate Law

For Runs 1-3, the initial I^- concentration is varied while the initial $\text{S}_2\text{O}_8^{2-}$ concentration remains constant.

The order with respect to I^- is determined using the method of initial rates. The following example will illustrate how to find a reaction order using the method of initial rates.

Example: The following data was obtained for the reaction: $\text{A} + \text{B} \rightarrow \text{C}$

Experiment	[A], M	[B], M	Rate (M/s)
1	0.030	0.10	1.30
2	0.030	0.10	1.80
3	0.030	0.25	11.25

The general rate law for this example is

$$\text{Rate} = k[\text{A}]^x[\text{B}]^y$$

Since [A] changes between Experiment 1 and 2, while [B] remains constant, the order for A is obtained by taking the ratio of the rates from these two experiments:

$$\frac{1.80 \text{ M/s}}{1.30 \text{ M/s}} = \frac{k(0.030)^x(0.10)^y}{k(0.020)^x(0.10)^y}$$

Since k is constant at a given temperature and $[\text{B}]^y$ is constant for Experiments 1 and 2, the equation simplifies to:

$$\frac{1.80 \text{ M/s}}{1.30 \text{ M/s}} = \left(\frac{0.030}{0.020}\right)^x \quad \text{or} \quad 1.30 = 1.5^x$$

Thus, $x = 1$ for this example.

Unfortunately, experimental results are not usually that "clean", and a more sophisticated method is needed to find x . Mathematically, solving for exponents requires the use of logarithms. Taking the log of both sides of the equation above yields:

$$\log(1.30) = x \cdot \log(1.5)$$

Rearranging this equation to solve for x yields

$$x = \frac{\log(1.30)}{\log(1.5)} = 1.0 = 1$$

Experiments 2 and 3 may then be used to find the order for B, as shown below:

$$\frac{11.25 \text{ M/s}}{1.80 \text{ M/s}} = \frac{k(0.030)^x(0.25)^y}{k(0.030)^x(0.10)^y}$$

By cancelling out the common terms and dividing the rate and concentration values, we obtain

$$6.25 = 2.5^y$$

Taking the log of both sides and rearranging to solve for y gives

$$y = \frac{\log(6.25)}{\log(2.5)} = 2.0 = 2$$

For Runs 2, 4, and 5, the initial $\text{S}_2\text{O}_8^{2-}$ concentration will vary while the initial I^- concentration remains constant. Use the initial rates method shown above to find the order for $\text{S}_2\text{O}_8^{2-}$.

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