

1. How many atoms are in 4 moles of iron?

$$4 \times 6.02 \times 10^{23} = 2.4 \times 10^{24}$$

2. What is the percentage by mass of oxygen in water?

$$\frac{16}{18} \times 100 = 89\%$$



$$H_2 = 2 \times 1 = 2$$

$$O = 1 \times 16 = 16$$

3. A compound is 57.5 % sodium, 40% oxygen and 2.5 % hydrogen. What is its empirical formula?

$$57.5g \times \frac{1 \text{ mol}}{23g \text{ Na}} = 2.5 \text{ mol Na}$$

$$2.5g \text{ H} \times \frac{1 \text{ mol}}{1g \text{ H}} = 2.5 \text{ mol H}$$

$$40g \times \frac{1 \text{ mol}}{16g \text{ O}} = 2.5 \text{ mol O}$$

$$1:1:1 \text{ ratio } \boxed{NaOH}$$

4. A compound with an empirical formula of CH_2 has a molecular weight of 42. What is its molecular formula?



$$C = 3 \times 12 = 36$$

$$H = 6 \times 1 = 6$$

$$42$$

5. At STP ~~48~~ 44.8 liters of neon would contain the same number of molecules as how many grams of Argon? What volume would the argon occupy?

$$44.8 \text{ L Ne} \times \frac{1 \text{ mol Ne}}{22.4 \text{ L Ne}} = 2 \text{ mol Ne}$$

$$2 \text{ mol Ar} \times \frac{39.9g \text{ Ar}}{1 \text{ mol Ar}} = 79.8g \text{ Ar}$$

$$\text{volume of Ar} = 44.8 \text{ L}$$

6. Sodium and chlorine react to form sodium chloride. If 4 moles of chlorine react, how many moles of sodium chloride will form?



$$4 \text{ mol } Cl_2 \times \frac{2 \text{ mol Na}}{1 \text{ mol } Cl_2} = 8 \text{ mol Na}$$

7. How many liters of chlorine at STP would be needed to react with 92 grams of sodium?

$$92g \text{ Na} \times \frac{1 \text{ mol Na}}{23g \text{ Na}} \times \frac{1 \text{ mol } Cl_2}{2 \text{ mol Na}} \times \frac{22.4 \text{ L}}{1 \text{ mol } Cl_2} = 44.8 \text{ L } Cl_2$$

8. How many milliliters of chlorine were required in question 7?

$$44.8 \text{ L} \times \frac{1000 \text{ mL}}{1 \text{ L}} = 44,800 \text{ mL } Cl_2$$

9. Write and balance the equation for the combustion of butane.



10. If 232.0 grams of butane are combined with 192 grams of oxygen and a reaction occurs, how many grams of water will form?

$$232g \text{ C}_4\text{H}_{10} \times \frac{1 \text{ mol C}_4\text{H}_{10}}{58g \text{ C}_4\text{H}_{10}} \times \frac{10 \text{ mol H}_2\text{O}}{2 \text{ mol C}_4\text{H}_{10}} \times \frac{18g \text{ H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 360g \text{ H}_2\text{O}$$

$$192g \text{ O}_2 \times \frac{1 \text{ mol O}_2}{32g \text{ O}_2} \times \frac{10 \text{ mol H}_2\text{O}}{13 \text{ mol O}_2} \times \frac{18g \text{ H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{83g \text{ H}_2\text{O}}$$

11. If the reaction in number 9 was performed and 2 liters of carbon dioxide were produced at STP, what was the percent yield?

$$2 \text{ L } CO_2 \times \frac{1 \text{ mol } CO_2}{22.4 \text{ L } CO_2} \times \frac{10 \text{ mol H}_2\text{O}}{8 \text{ mol } CO_2} \times \frac{18g \text{ H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 2g \text{ H}_2\text{O}$$

$$\frac{2g \text{ H}_2\text{O}}{83g \text{ H}_2\text{O}} \times 100 = \boxed{2.4\%}$$

12. Write and balance the equation for the reaction of aluminum with hydrochloric acid.



$$\frac{2g \text{ H}_2}{2g \text{ H}_2} \times \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} \times \frac{27g \text{ Al}}{1 \text{ mol}} = 13.5g \text{ Al}$$

13. If you are given 108 grams of aluminum to react with 146 grams of HCl, what will be the limiting reactant?

$$108g \text{ Al} \times \frac{1 \text{ mol Al}}{26.98g \text{ Al}} \times \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} \times \frac{2g \text{ H}_2}{1 \text{ mol H}_2} = 12g \text{ H}_2$$

$$146g \text{ HCl} \times \frac{1 \text{ mol HCl}}{36.45g \text{ HCl}} \times \frac{3 \text{ mol H}_2}{6 \text{ mol HCl}} \times \frac{2g \text{ H}_2}{1 \text{ mol H}_2} = \boxed{4g \text{ H}_2}$$

HCl is limiting reactant

Chemistry Semester 2 Practice Final

**Jeanne Mahoney Leffers, Jennifer Gail
Audette, Kevin S. Hardwick, William
Van Cleve**

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Register University of California, Berkeley, 1920 **General Catalog** University of Missouri, 1892 **Progressive Age**, 1908 **Bulletin of Clarkson College of Technology** Clarkson College of Technology, 1912 **Catalogue of the Officers and Students of Antioch College of Yellow Springs, Greene Co., Ohio, for the Academical Year** Antioch College, 1903 International Partnerships for Strengthening Health Care Workforce Capacity: Models of Collaborative Education Jeanne Mahoney Leffers, Jennifer Gail Audette, Kevin S. Hardwick, William Van Cleve, 2018-07-31

A critical problem in resource scarce countries across the globe is the shortage of appropriately trained health care providers. According to the World Health Organization the current global health workforce shortage of 7.2 million providers is estimated to increase to 12.9 million by 2035. This disproportionately affects resource scarce countries denying basic health care to millions and limiting access to life saving treatments. Due to limited resources in these countries not enough health professionals receive training few have the opportunity for continuing education and the ability to develop or implement educational programs and curricula is constrained. Additionally many existing providers choose to emigrate in pursuit of professional advancement opportunities contributing to the overall shortage of qualified health care providers in these environments. Efforts to strengthen health workforce capacity not only increases access safety and availability of care but is critical to building resilient health systems capable of caring for the world's neediest populations. This requires not only cultivating new health care providers but also providing ongoing professional development to retain and support current providers advancing the level of practice in accordance with current clinical science cultivating educators and enhancing training curricula. It is critical also to contribute to the limited body of research documenting the effectiveness and impact of various models of collaborative education and partnership to improve health worker training and retention. This Research Topic examines strategies for building health workforce capacity through the prism of educational partnerships offering significant examples of effective models of international collaborative education as well as insight and guidance on the structure and operation of successful global partnerships. Collectively the 31 articles accepted and included in this eBook represent a diversity of health professions and geographies across academic non governmental organizations and other global partnership forms. The published manuscripts highlight various elements of partnerships with several consistent themes emerging capacity building local empowerment mutual trust and respect long term commitment equity collaboration and the importance of integrating theory and practice for a balance of academic and clinical development. The manuscripts provide examples of partnership and educational programs that are in the formative early stages of implementation and others which have been sustained long term some for decades. The following eBook is divided into two parts with each part broken down into sections. Part I of the eBook includes 18 manuscripts that showcase long term educational programs that strongly exemplify multiple foundational aspects of international partnerships in education including mutual collaboration and project management empowerment of

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