## 11-2 Practice Problems

 Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:

Eglin x 1 mai Liju x 3 mai Liot x 24 g Liot = 0.78 g 1 x 35 g Liju x 1 mai Liot x 1 mai Liot Liot Liot

> What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?

029 NoI + Cl2 - 2NoCI + Iz 029 NoI + Cl2 - 2NoCI + Iz 029 NoI x 1mol x 2mol NoI (mol NoC) NoCI

> Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:

0+85gCyHo x 1mol x 5mol Cox x 44g = 2.58g 1 x 58g x 2mol CyHo 1mol CO z

 Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:

0.4695bD= x [mol x 2mol 5b, 03 x 121.769 fa.389 Sb

5. What mass of hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>)
must decompose to produce 0.77 g of
water?
2H<sub>2</sub>O<sub>2</sub> -> 2H<sub>2</sub>O + O<sub>2</sub>

0.779 H20 x (mol x 2420 x 349 = [.4594:02

6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?

0.69gCOz × 1mol × 2mol CO x 289 = 0.449 1 449 2mol COz 1mol = 0.449  Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:

 $\frac{N8(NO_3)_2 + 2N8OH \rightarrow N8(OH)_2 + 2N8NO_3}{O.73_3 V80_3)_2 \times \frac{1000}{185_3} \times \frac{2001 N8NO_3}{1001 N6NO_3)_2} \times \frac{85_3}{1001} = \frac{(2.68_3)}{1001 N6NO_3}$ 

 Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:

 $0.64_9H_2O \times \frac{CaC_2 + 2H_2O - Ca(OH)_2 + C_2H_2}{189} \times \frac{1001(a(0H)_2 + C_2H_2)}{1001} \times \frac{749}{1001} = \frac{1.329}{Ca(0H)_2}$ 

How many grams of ozone (O<sub>3</sub>) must decompose to produce 0.87 g of oxygen?

0.8790 x 1mg/ x 2mg/ x 48903 = 0.87909

 Find the mass of sugar (C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation: C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> → 2C<sub>2</sub>H<sub>6</sub>O + 2CO<sub>2</sub>

 How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP?
 The balanced equation is S + O<sub>2</sub> → SO<sub>2</sub>.

 Find the mass of benzene (C<sub>6</sub>H<sub>6</sub>) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:

2C4H6 + 15O2 → 6H2O + 12CO2

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**Walter Kauzmann** 

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Part One features two introductory chapters intended to introduce the mutual interests of engineers chemists geologists and industrial scientists in the physical and chemical properties of oxide materials Part Two provides the essential and fundamental principles that are critical to understanding most of the major reactions between water and oxides Part Three deals with the synthesis of oxide materials in aqueous media Part Four deals with oxide water reactions and their environmental and technological impacts and Part Five is devoted to other types of relevant reactions The Aqueous Chemistry of Oxides is the first book that provides a comprehensive summary of all of the critical reactions between oxides and water in a single volume As such it ties together a wide range of existing books and literature into a central location that provides a key reference for understanding and accessing a broad range of more specialized topics The book contain over 300 figures and tables Study Guide and Solutions Manual for Organic Chemistry Susan McMurry, John McMurry, 1988

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