

Chemistry

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(Chapter - 3) (Chemical Kinetics) (Exercises)
(Class 12)

Question 3.1:

From the rate expression for the following reactions, determine their order of reaction and the dimensions of the rate constants.



Answer 3.1:



Therefore, order of the reaction = 2

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{NO}]^2} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})^2} = \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol}^2\text{L}^{-2}} = \text{L mol}^{-1}\text{s}^{-1}$$



Therefore, order of the reaction = 2

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{H}_2\text{O}_2][\text{I}^-]} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})(\text{mol L}^{-1})} = \text{L mol}^{-1}\text{s}^{-1}$$



Therefore, order of reaction = $3/2$

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{CH}_3\text{CHO}]^{3/2}} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})^{3/2}} = \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol}^{3/2}\text{L}^{3/2}} = \text{L}^{3/2}\text{mol}^{1/2}\text{s}^{-1}$$



Therefore, order of the reaction = 1

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{C}_2\text{H}_5\text{Cl}]} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol L}^{-1}} = \text{s}^{-1}$$

Question 3.2:

For the reaction:



The rate = $k[\text{A}][\text{B}]^2$ with $k = 2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$. Calculate the initial rate of reaction when $[\text{A}] = 0.1 \text{ mol L}^{-1}$, $[\text{B}] = 0.2 \text{ mol L}^{-1}$. Calculate the rate of reaction after $[\text{A}]$ is reduced to 0.06 mol L^{-1} .

Answer 3.2:

The initial rate of reaction is

$$\text{Rate} = k[\text{A}][\text{B}]^2 \Rightarrow (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1})(0.1 \text{ mol L}^{-1})(0.2 \text{ mol L}^{-1})^2 = 8.0 \times 10^{-8} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$$

When $[\text{A}]$ is reduced from 0.1 mol L^{-1} to 0.06 mol L^{-1} ,

$$\text{The concentration of A reacted} = (0.1 - 0.06) \text{ mol L}^{-1} = 0.04 \text{ mol L}^{-1}$$

$$\text{Therefore, concentration of B reacted} = \frac{1}{2} \times 0.04 \text{ mol L}^{-1}$$

$$\text{Then concentration of B available, } [\text{B}] = (0.2 - 0.02) \text{ mol L}^{-1} = 0.18 \text{ mol L}^{-1}$$

After $[\text{A}]$ is reduced to 0.06 mol L^{-1} , the rate of reaction is given by,

$$\text{Rate} = k[\text{A}][\text{B}]^2 \Rightarrow (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1})(0.06 \text{ mol L}^{-1})(0.18 \text{ mol L}^{-1})^2 = 3.89 \text{ mol L}^{-1} \text{ s}^{-1}$$

Question 3.3:

The decomposition of NH_3 on platinum is zero order reaction. What are the rates of production of N_2 and H_2 if $k = 2.5 \times 10^{-4} \text{ mol}^{-1} \text{ L s}^{-1}$?

Answer 3.3:

The decomposition of NH_3 on platinum surface is represented by the following equation.



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Solutions Manual for Perspectives on Structure and Mechanism in Organic Chemistry Felix A. Carroll, 2023-05-02
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Superoxide Ion Chemistry and Biological Implications Igor B. Afanas'ev, 2024-12-11 The chemical properties of superoxide ion its biological role and the role of other oxygen radicals which arise as a result of its transformations are contained in this text In Volume I the principal reactions of superoxide ion including protonation reactions with proton donors nucleophilic reactions with esters alkyl halides and other compounds electron transfer reactions with quinones and metal complexes are described Basic quantitative data including rate constants and yields for the reactions of superoxide ion of all types are given in tables This volume contains the mechanisms of the generation of oxygen radicals in cells and the interaction of superoxide ion with cell components The role of superoxide ion in lipid peroxidation and destruction of proteins and nucleic acids is explained as well as oxygen radicals in the mechanisms of toxic and therapeutic action of drugs especially anticancer antibiotics In addition the action of superoxide ion and other oxygen radicals on plants micro and macroorganisms is discussed along with the role of oxygen radicals in normal metabolic and pathological processes

Free-Radical-Induced DNA Damage and Its Repair Clemens Sonntag, 2006-03-20 The free radical chemistry of DNA had been discussed in some detail in 1987 in my book *The Chemical Basis of Radiation Biology* Obviously the more recent developments and the concomitant higher level of understanding of mechanistic details are missing Moreover in the living cell free radical DNA damage is not only induced by ionizing radiation but free radical induced DNA damage is a much more general phenomenon It was therefore felt that it is now timely to review our present knowledge of free radical induced DNA damage induced by all conceivable free radical generating sources Originally it had been thought to include also a very important aspect the repair of DNA damage by the cell's various repair enzymes Kevin Prise Cancer Campaign Gray Laboratory London was so kind to agree to write this part However an adequate description of this strongly expanding area would have exceeded the allocated space by much and this section had to be omitted The directors of the Max Planck Institut für Strahlenchemie now MPI für Bioanorganische Chemie Karl Wieghardt and Wolfgang Lubitz kindly allowed me to continue to use its facilities after my retirement in 2001 Notably our librarian Mrs Jutta Theurich and her right hand help Mrs Rosemarie Schröder were most helpful in getting hold of the literature I thank them very much Without their constant help this

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