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CHEMISTRY: COUNTING ATOMS IN COMPOUNDS WORKSHEET #7.0.1

INSTRUCTIONS: Write the quantity of atoms of each element opposite the formula of the compound for the quantity of formula units and molecules shown:

For example: $5P_2O_5$ $P = (5 \times 2 =) 10$ $O = (5 \times 3 =) 15$

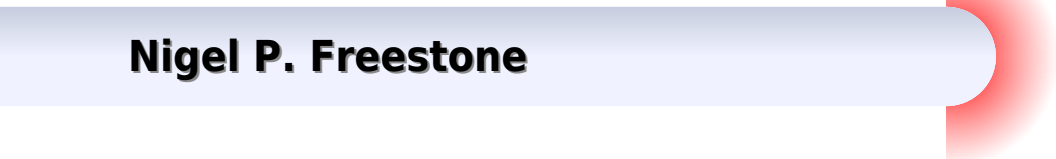
For example: $4Zn(NO_3)_2$ $Zn = (4 \times 1 =) 4$ $N = (4 \times 1 \times 2 =) 8$ $O = (4 \times 3 \times 2 =) 24$

- | | | | | |
|-----|--------------------|--------------|--------------|-------------------------|
| 1. | $4K_2CO_3$ | $K =$ _____ | $C =$ _____ | $O =$ _____ |
| 2. | $2Sr_3(PO_4)_2$ | $Sr =$ _____ | $P =$ _____ | $O =$ _____ |
| 3. | $3N_4O_{10}$ | $N =$ _____ | $O =$ _____ | |
| 4. | $2(NH_4)_3N$ | $N =$ _____ | $H =$ _____ | |
| 5. | $8Cl_2O$ | $Cl =$ _____ | $O =$ _____ | |
| 6. | $Ca(C_2H_3O_2)_2$ | $Ca =$ _____ | $C =$ _____ | $H =$ _____ $O =$ _____ |
| 7. | $12NaBr$ | $Na =$ _____ | $Br =$ _____ | |
| 8. | $4Al(OH)_3$ | $Al =$ _____ | $O =$ _____ | $H =$ _____ |
| 9. | $3NaHCO_3$ | $Na =$ _____ | $H =$ _____ | $C =$ _____ $O =$ _____ |
| 10. | $5Ga_2(Cr_2O_7)_3$ | $Ga =$ _____ | $Cr =$ _____ | $O =$ _____ |
| 11. | $7C_2S_2$ | $C =$ _____ | $S =$ _____ | |
| 12. | $4Fe_2O_3$ | $Fe =$ _____ | $O =$ _____ | |
| 13. | $6Ba(MnO_4)_2$ | $Ba =$ _____ | $Mn =$ _____ | $O =$ _____ |
| 14. | $3V_2O_5$ | $V =$ _____ | $O =$ _____ | |
| 15. | $2KNO_3$ | $K =$ _____ | $N =$ _____ | $O =$ _____ |
| 16. | $9MgSO_4$ | $Mg =$ _____ | $S =$ _____ | $O =$ _____ |
| 17. | $5Al_2(SiO_3)_2$ | $Al =$ _____ | $Si =$ _____ | $O =$ _____ |
| 18. | $4Au(IO_3)_3$ | $Au =$ _____ | $I =$ _____ | $O =$ _____ |

(Continued)

Chemistry Counting Atoms In Compounds

Nigel P. Freestone



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