

Chemistry Common Exam Study Guide and Checklist

Objective 1 - Matter: Properties and Change

Chem. 1.1 Big Idea: Analyze the Structure of Atoms and Ions.

(9 Instructional Days)

- ❖ *Analyze the structure of atoms, isotopes, and ions.*
- ❖ *Analyze an atom in terms of the location of electrons.*
- ❖ *Explain the emission of electromagnetic radiation in spectral form in terms of the Bohr model.*
- ❖ *Explain the process of radioactive decay using nuclear equations and half-life.*

Objective: 1.1.1

- ☐ Characterize protons, neutrons, electrons by location, relative charge, relative mass ($p=1$, $n=1$, $e=1/2000$).
- ☐ Use symbols: A = mass number, Z =atomic number
- ☐ Use notation for writing isotope $\text{U} \begin{smallmatrix} 235 \\ 92 \end{smallmatrix}$ or U-235
- ☐ symbols:
- ☐ Identify isotope using mass number and atomic number and relate to number of protons, neutrons and electrons.
- ☐ Differentiate average atomic mass of an element from the actual isotopic mass and mass number of specific isotopes. (Use example calculations to determine average atomic mass of atoms from relative abundance and actual isotopic mass to develop understanding).

Objective: 1.1.2

- ☐ Analyze diagrams related to the Bohr model of the hydrogen atom in terms of allowed, discrete energy levels in the emission spectrum.
- ☐ Describe the electron cloud of the atom in terms of a probability model.
- ☐ Relate the electron configurations of atoms to the Bohr and electron cloud models.

Objective: 1.1.3

- ☐ Understand that energy exists in discrete units called quanta.
- ☐ Describe the concepts of excited and ground state of electrons in the atom:
 1. Gaining energy results in the electron moving from its ground state to a higher energy level.
 2. When the electron moves to a lower energy level, it releases the energy difference in the two levels as electromagnetic radiation (emissions spectrum).

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