



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol $\rightleftharpoons$.

What is Equilibrium? Chemical Equilibrium is a state where the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol $\rightleftharpoons$. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant (K_p) For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	
Equilibrium Constant, K_c For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant, K_p For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	
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Le Chatelier's Principle: the effect of changing conditions on equilibrium		
Concentration	Increasing concentration of a reactant will shift the equilibrium to the right. Increasing concentration of a product will shift the equilibrium to the left.	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the concentration of N_2 will shift the equilibrium to the right, producing more NH_3.
Pressure	Increasing pressure will shift the equilibrium to the side with the fewest moles of gas. Decreasing pressure will shift the equilibrium to the side with the most moles of gas.	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the pressure will shift the equilibrium to the right, producing more NH_3.
Temperature	Increasing temperature will shift the equilibrium to the endothermic side. Decreasing temperature will shift the equilibrium to the exothermic side.	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the temperature will shift the equilibrium to the left, producing more N_2 and H_2.

Chemistry Chemical Kinetics And Equilibrium Study Guide

Victor M. Corman



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