

Acids and Bases

Overview of Acids and Bases

Indicators: Acids and Bases

Acids: Donates its proton to base (proton donor)

Base: Donates its lone pair to acid (LBP)



Remember: (1) only H atoms, (2) only OH is a base

What about organic indicators and indicators solution?

Acids and Bases

Acids are proton donors, (the electrons is being to move)
Bases are proton acceptors, (the electrons is still being the base)



General acid-base reactions of acid-base reactions:

to find the strong base, the base gets negative charge of the base

and check for weak acids hydrogens of the acid.

to find the level of the hydrogens and move the electrons

to the other atom

Remember: -OH is not always base, depends on the context

1. Acid: Donates and Proton: Donates, are from Proton: Protonates

What if there is no proton transfer?

Acids and Bases

Acids: acid is an electron-pair acceptor.

Acids: base is an electron-pair donor.



Acid-base reactions: acids and bases are described in terms of electron.

1. One of the acid-base reactions: transfer of one electron

As a general reaction: transfer of one electron to (1) and from (2)

Acid Strength

What does it mean "strong acid"?

Relatively speaking, the more the acid donates its proton, the stronger it is.



The dissociation of the acid is a reversible reaction that is described by equilibrium constant



Acid Strength and pK_a

The K_a values of weak acids (1) and weak bases (2) are shown below



The acid is a stronger acid than the base

To compare the acids, pK_a is used, which is the pK_a of the given acid

$$pK_a = -\log K_a$$

$$pK_a = -\log K_a = 10 \Rightarrow pK_a = -\log K_a = 10$$

So, the stronger the acid, the lower the pK_a value (It is to separate the two)

Acids are pK_a values of weak acids (strong acids)



What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base (the base)

The stronger the acid, the more stable the conjugate base.



Acids (1) acids are weak acids, bases are strong acids

There are a few factors to be considered when assessing the stability of the conjugate base.

1. The first and the last factor is the **charge** to which the hydrogens is connected. Usually

to **anion** (for hydroxides and hydroxides).

2. When the acid is strong, then the charge is to be strong, the more



Stronger acids: stronger acids are stronger acids, the more stable

$$pK_a = 10 \Rightarrow pK_a = 10$$

1. When comparing charges in different ways, the more stable is the base

Large acids have higher polarizability and results for negative charge better in the

charge is spread out



There are weak acids and strong acids in weak acids, the more stable

The more stable is the base, the more stable the negative charge better

2. When comparing for same atom, check which conjugate base is **stronger** (stronger)



1. The first factor for determining the acidity is called **inductive**

Inductive effect is the transmission of charge through atoms

Electronegativity atoms pull the negative charge and help to stabilize the conjugate base



The more electronegative, the more electron withdrawing, the stronger the acid



The closer the electronegative atom to the hydrogens, the stronger the acid

Increasing Acidity

2. The second factor is the **resonance** of the acid, which is the hydrogens

The explains the pK_a values of acids, bases, and bases



3. Which is more electronegative than F, or the more F character, the stronger the acid

Therefore, for **acids**, **inductive** > **resonance** > **charge**

The primary acids: **Acids**, **Bases**, **Inductive**, **Resonance**, **Charge**

of a compound is called as **acidity**

Remember: the stability of the conjugate base



The Position of Equilibrium for an Acid-Base Reaction

Is the following reaction reaction possible?



Acid-base reactions are reversible, so neither way of doing the reaction is

to the equilibrium of the reaction, which is the right (products) or the left (reactants)?

To answer this question, you need to consider: for the equilibrium of acids and bases

reaction is shifted towards the weaker acid and the base

So, you need to compare the pK_a of the acid and the pK_a of the conjugate acid

The equilibrium will be shifted to the higher pK_a reaction with only



Is the acid the pK_a of the conjugate acid is greater than the acid pK_a , or the

equilibrium is shifted to the right. Therefore, the right reaction is more likely

Choosing a Base to Deprotonate a Given Compound

Suppose you need to deprotonate the following compound: **acid**

Deprotonate means remove the proton from the acid

First, find the pK_a value: **acid** $pK_a = 10$

Find the chemical equation of the conjugate base with a conjugate base (1)



The pK_a of the conjugate acid must be higher than 10, so the reaction will be favored

Find a conjugate base with $pK_a > 10$. Acids are weak acids $pK_a > 10$



Choosing an Acid to Protonate a Given Compound

Suppose you need to protonate the following compound: **acid** (1) (2) (3) (4) (5)

Which is a very reactive base and needs to be protonated before being disposed

To do that, write the chemical equation of acid + base pK_a



The pK_a of the acid (the conjugate acid of acid) is 10. Therefore, the pK_a of **acid**

and must be smaller than 10, so the reaction will be favored

Acid + base and it is a proton donor (acid) for protonating the acid + base

in $pK_a > 10$ and is a base stronger than the acid base



Hammett-Hansch Equation

The equation shows whether the compound with a given pK_a is to be acidic or

base (deprotonated) base is an organic solution depending on the pK_a of the

acid

$$pK_a = pK_a + \log \frac{[A^-]}{[HA]}$$

One general and application of the Hammett-Hansch equation

1. If the $pK_a > pK_a$ 2. If the $pK_a < pK_a$ 3. If the $pK_a = pK_a$

$$\Delta \log \frac{[A^-]}{[HA]} = 0$$

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James E. Brady, Fred Senese



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