



Reaction Rates and Equilibrium Study Guide

ANSWER KEY

- Describe how each of the following factors would affect the rate of a chemical reaction in terms of "particle collisions" in your response.
 - Temperature decreases** *If temperature decreases, the particles move slower and there will be less reactant particle collisions AND a smaller percentage of particles will have enough energy to produce a reaction. The reaction rate will therefore decrease.*
 - Addition of iodide ions, I^- , as a catalyst** *The presence of the catalyst lowers the activation energy of the reaction taking place. This means that a larger percentage of particles will have enough energy to produce a reaction. The reaction rate will therefore increase.*
 - Increased reactant concentration** *With more reactant particles present, there are more collisions between those particles. This means the reaction rate will increase.*
 - Crushing a solid into a powder (with larger surface area)** *As a powder, there is more surface area for collisions to take place. This means there will be more collisions and therefore a faster reaction rate.*
- The reaction progress diagram shows the energy changes that take place when nitrogen dioxide gas reacts with carbon monoxide gas to form carbon dioxide gas and nitrogen gas.

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Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students **Organic Chemistry, 4e**

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