

Chemical Reactions

Chemical Equation

Reactants $\longrightarrow$ Products
Left hand side Right hand side

magnesium + Oxygen $\longrightarrow$ magnesium oxide



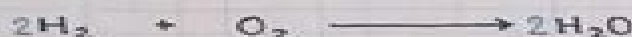
BALANCING EQUATIONS



Chemical equations are a way of showing chemical reactions.

LAW OF CONSERVATION OF MATTER
a chemical equation must be balanced.

number of atoms of every element before and after reaction is same



Chemical reactions

COMBINATION

Two or more substance combine to form a single product.



DECOMPOSITION

Single reactant breaks up to form two or more products



DISPLACEMENT

More reactive element removes the other element from the compound and itself forms the compound

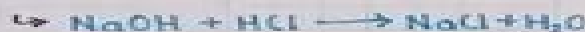


- Chemical reactions can occur only:
 $\hookrightarrow$ when the chemical change.

Rusting of Iron, Burning of coal.

DOUBLE DISPLACEMENT

Exchange of ions between the reactants take place.



REDUCTION

gain of hydrogen or loss of oxygen.

redox

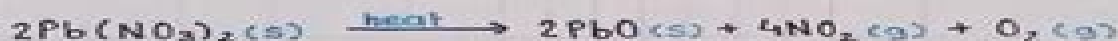
OXIDATION

gain of oxygen or loss of hydrogen.

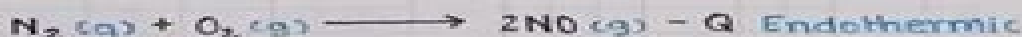


reduction

PHYSICAL STATE -



CHANGE IN HEAT -



EVOLUTION OF GAS -



PRECIPITATION -



DIFF. REACTIONS CONDITION -



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Chemistry Notetaking Guide

Scott Zimmerman



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