

Chemistry Note Taking Guide

Describing Matter (p.6-15)

- *Matter* is anything that has mass and takes up space (NRG is everything else)
- Chemistry is the study of the properties of matter & how matter Δ
- *Substance* - single type of matter that is pure; specific make-up; specific properties (salt vs. blueberry muffin)
- Chemical & Physical properties describe matter
- *Physical Property* - can be observed without changing it into another substance (water to ice or water to steam); hardness, texture, color, state of matter; can be used to classify matter; usually observed with the five senses
- *Chemical Property* - characteristic of a substance that describes its ability to Δ into a different substance; flammability, rusting, tarnishing, production of gas
- *Elements* - 117 known ones, simplest substances; can't be broken down into any other substance (must be 118 - haven't made element #117 yet)
- Chemical bonds form when atoms combine to make *molecules* - groups of two or more atoms held together by chemical bonds
- *Compound* is a pure substance made of two or more elements chemically combined in a set ratio (comparison of two numbers); have properties that are different from those of the uncombined elements; difficult to separate into its elements; have a chemical formula that can be written
- Compounds are represented by chemical formulas - shows the elements and the ratio of atoms of those elements (H_2O ; CO_2 ; $NaCl$; $C_6H_{12}O_6$)
- Both elements and compounds are "pure" substances
- *Mixture* is made from two or more substances that are NOT chemically combined; easy to separate into its components (filtration, evaporation, distillation, etc), cannot be written as ONE chemical formula
- *Heterogeneous Mixture* - not evenly mixed; can see the different parts (salad)
- *Homogeneous Mixture* - evenly mixed; can't see the different parts (sugar + water)
- *Solution* - example of a homogeneous mixture - could be solid (brass - copper + zinc), liquid (lemonade) or gas (air)

Measuring Matter (p. 16-20)

- Unlike weight, mass does not change with location, even when the force of gravity changes on an object.
- *International System of Units* - (SI) - used to measure the properties of matter - metric
- Mass, Weight, Volume, Temperature, and Density are physical properties of matter.
- Density of water = 1 gm/cm^3

States of Matter (p. 42 - 47)

Chemistry Note Taking Guide 15

Carl Friedrich PESCHEL



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The Complete Idiot's Guide to Study Skills Randall S. Hansen, 2008-09-02 Study smarter not harder Every high school and college student would love to know how to get the highest grades with the least amount of effort This book gives students a guiding philosophy for every class every time laying the foundation for lifelong learning With the wisdom gained from these tips success stories from other students and mini assessments they ll be empowered to succeed in class preparation reading comprehension exam taking and more No one method fits every student so included are many tried and true methods Useful for every subject from foreign languages to mathematics from high school through college and beyond Helps students find their particular learning styles

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