

COMPOSITION OF HYDRATES

Topic: Hydrates

A hydrate is an ionic compound with water molecules loosely bonded to its crystal structure. The number of water molecules bonded to each unit of the salt. For example, the formula $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ indicates that there are x water molecules for every one formula unit of CuSO_4 . Answer the questions below.

1. What percentage of water is found in $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$?

$$C_{\text{Cu}} = 64$$

$$S_{\text{S}} = 32$$

$$O_{\text{O}} = 16$$

$$H_{\text{H}} = 1 \quad \text{Total} = 159.5$$

$$20 \text{ H}_2\text{O} = \frac{360}{159.5} = 2.26\%$$

$$36.7\%$$

2. What percentage of water is found in $\text{FeSO}_4 \cdot x\text{H}_2\text{O}$?

$$\text{Fe} = 56 \quad S_{\text{S}} = 32 \quad O_{\text{O}} = 16$$

$$H_{\text{H}} = 1 \quad \text{Total} = 159.5$$

$$18 \text{ H}_2\text{O} = \frac{324}{159.5} = 2.03\%$$

$$\frac{18 \times 18}{159.5} \times 100 = 20.3\%$$

$$6.57\%$$

3. A 5.0 gram sample of a hydrate of BaCl_2 was heated, resulting in 3.0 g of the anhydrous salt remaining. What percentage of water was in the hydrate?

$$\text{Ba} = 137 \quad \text{Cl} = 35.5 \quad \text{H}_2\text{O} = 18$$

$$\text{Hydrate} = \text{mass} \quad \text{Anhydrous} = \text{mass}$$

$$\text{Water} = \text{mass} \quad \text{Anhydrous} = \text{mass} \quad \text{Hydrate} = \text{mass} \quad \text{Anhydrous} = \text{mass}$$

$$\frac{2.0}{5.0} \times 100 = 40\%$$

$$14.7\%$$

4. A 2.5 g sample of a hydrate of $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ was heated, and only 1.7 g of the anhydrous salt remained. What percentage of water was in the hydrate?

$$\frac{1.7}{2.5} \times 100 = 68\%$$

$$\frac{1.7}{2.5} \times 100 = 68\%$$

$$36.7\%$$

5. A 3.0 gram sample of $\text{Na}_2\text{CO}_3 \cdot x\text{H}_2\text{O}$ is heated to constant mass. How much anhydrous salt remains?

$$2 \text{ Na} = 46 \quad \text{C} = 12 \quad \text{O}_3 = 48 \quad \text{H}_2\text{O} = 18$$

$$\frac{18}{112} \times 100 = 16.07\%$$

$$\frac{18}{112} \times 100 = 16.07\%$$

$$\frac{18}{112} \times 100 = 16.07\%$$

$$0.49\%$$

6. A 5.0 gram sample of $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ was heated, and 3.0 g of the anhydrous salt remained. Determine the value of x .

$$\text{Cu} = 64 \quad \text{S} = 32 \quad \text{O}_4 = 64 \quad \text{H}_2\text{O} = 18$$

$$\frac{159.5}{179.5} \times 100 = 88.8\%$$

$$\frac{159.5}{179.5} \times 100 = 88.8\%$$

$$\text{CuSO}_4 = 159.5$$

$$18x = 20$$

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