-NO ₃ ² all soluble		-OH all insoluble except Group 1, Ba(OH)z, Sr(OH)z, Ca(OH)z		
-Cl' all soluble except AgCl, Hg ₂ Cl ₂ and PbCl ₂		-CO ₅ ² all insoluble except NH ₄ * and Group 1		
50 ₄ ² all soluble excep	ot BaSO ₄ , PbSO ₄ and SrSO ₄	-S ⁻² all insoluble	except Group 1	and Group 2 and NH _s *
Polyatomics:				Acids:
Acetate (C ₂ H ₂ O ₂ ⁻¹)	-Carbonate (CO ₃ ²)	-Arsenate (AsO	3	-HCI = hydrochloric
Chlorate (CIO ₃ ⁻¹)	-Chromate (CrO ₄ ²)	-Arsenite (AsO ₃	")	-H ₂ SO ₄ = sulfuric
Chlorite (CIO ₂ ⁻¹)	-Dichromate (Cr ₂ O ₇ -2)	-Phosphate (PO	• ")	-HNO ₃ = nitric
Cyanide (CN ⁻³)	-Oxalate (C ₂ O ₄ ²)	-Phosphite (POs	*)	-CH ₂ COOH = acetic
Hydroxide (OH1)	-Sulfate (SO _{v-2})	-Ammonium (Ni	Ha*1)	-H ₃ PO ₄ = phosphoric
Nitrate (NO ₃ ¹)	-Sulfite (SO ₃ ²)	-Bicarbonate (H	CO ₃ ⁽³)	-H ₂ CO ₃ = carbonic
Nitrite (NO ₂ °)				
Nitrite (NO ₂ °)				
Matter is the physical i	material of the universe and	exists in 3 states: solid, liqu	iid, gas	
iubstance is matter th	at has a uniform and definite	e composition	Mixture is a p	hysical blend of substances
lomogenous Mixture	has uniform composition the	roughout (solution) Hetero	geneous Mixtur	e does not
Physical Properties: observed or measured		Chemical Properties: ability to undergo chemical change		
aw of Conservation of	f Mass: matter/mass can be	neither created nor destroy	red	
Extensive Properties: n	mass and volume	Intensive Properties: composition		
iolid: definite shape, v	volume, not easily compresse	ed <u>Liquid</u> : indefinite shape,	definite volume	, not easily compressed
	volume, not easily compressed volume, easily compressed			, not easily compressed
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Gas: indefinite shape, on the state of the shape, of the state of the	volume, easily compressed	*Particles are always mo	ving in S, L, G duction of gas,	
Gas: indefinite shape, on indications of Chemica Qualitative give descrip	volume, easily compressed il Reaction: dissolving, releas	*Particles are always mo sing/absorption of heat, pro Quantitative number-bas	ving in S, L, G duction of gas, s	
Gas: indefinite shape, on indications of Chemica Qualitative give descrip accuracy is how close to	volume, easily compressed il Reaction: dissolving, releas ptions, no numbers	*Particles are always mo ling/absorption of heat, pro Quantitative number-base Precision is how	ving in S, L, G duction of gas, o sed close several m	color change, precipitate neasurements are together
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Gas: indefinite shape, on indications of Chemica Qualitative give descrip accuracy is how close to firror = expt - accepts 1 kilo(unit) = 1000 (unit	volume, easily compressed il Reaction: dissolving, releas ptions, no numbers to a single measurement ed % Error = expt-accepte it) 1 (unit) = 1000 milli	*Particles are always mo sing/absorption of heat, pro Quantitative number-bas Precision is how di/accepted * 100 Percen 1 (unit) = 100 centi	ving in S, L, G duction of gas, s sed close several m t Yield = expt/th micro=10 ⁶	color change, precipitate neasurements are together neoretical * 100
Gas: indefinite shape, indications of Chemica Qualitative give descrip Accuracy is how close to Error = expt - accepte I kilo(unit) = 1000 (unit) C = C + 273	volume, easily compressed If Reaction: dissolving, release ptions, no numbers to a single measurement ed % Error = expt-accepte it) 1 (unit) = 1000 milli 1 - 273 H ₂ O FP=273K	*Particles are always mo ling/absorption of heat, pro Quantitative number-base Precision is how d]/accepted * 100 Percen 1 (unit) = 100 centi BP=373XDensity = mass/	ving in S, L, G duction of gas, o sed close several m t Yield = expt/th micro=10 ⁶ volume	neasurements are together neoretical * 100 nano=10 ^p
Gas: indefinite shape, on indications of Chemica Cauditations of Chemica Cauditation give descriptor accepts in the company is how close the company is shown close the company in the com	volume, easily compressed If Reaction: dissolving, release ptions, no numbers to a single measurement ed % Error = expt-accepte it) 1 (unit) = 1000 milli 1 - 273 H ₂ O FP=273K	*Particles are always mo sing/absorption of heat, pro Quantitative number-bar Precision is how d]/accepted * 100 Percen 1 (unit) = 100 centi BP=373XDensity = mass/ ecimal places / For M	ving in S, L, G duction of gas, o sed close several m t Yield = expt/th micro=10 ⁶ volume fultiplication & 6	color change, precipitate neasurements are together neoretical * 100 nano=10° Density of H ₂ O = 1 g/mL

Neutron = neutral, change = isotope

Electron = neg., change = ion

Proton = pos., change = radioactivity

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Yan Bai

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