

In your textbook, read about molarity and preparing molar solutions.

Read the following problem and then answer the questions.

An 85.0-mL aqueous solution contains 7.54 g iron(II) chloride (FeCl_2). Calculate the molarity of the solution.

5. What is the mass of the solute?

7.54 g

6. What is the volume of the solution?

85.0 mL solution

7. Write the equation that is used to calculate molarity.

$$M = \frac{\text{moles of solute}}{\text{L solution}}$$

8. In what unit must the amount of the solute be expressed to calculate molarity?

moles

9. In what unit must the volume of the solution be expressed to calculate molarity?

Liter

10. Write the expression needed to convert the volume of the solution given in the problem to the volume needed to calculate molarity.

$$85.0 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 0.0850 \text{ L}$$

Lesson 13.2 continued

11. What quantity must be used to convert the mass of the solute given in the problem to the amount calculate molarity?

molar mass of FeCl_2 126.75 g/mol

12. Write the expression used to calculate the amount of solute

$$7.54 \text{ g FeCl}_2 \times \frac{1 \text{ mol}}{126.75 \text{ g}} = 0.0595 \text{ mol}$$

13. Calculate the molarity of the solution. Show all your work.

$$\underbrace{7.54 \text{ g} \times \frac{1 \text{ mol}}{126.75}}_{\text{moles of Solute}} \times \underbrace{\frac{1}{0.0850 \text{ L}}}_{\text{L of Solution}} = \textcircled{0.700 \text{ M}}$$

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Chemistry Chapter 13 Study Guide

David R. Klein



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aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us *Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **Study Guide and Solutions Manual** Neil E. Schore, 2002-08-02 *Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition* Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

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