Chapter 21

Organic Chemistry: Some Basic Principles and Techniques

- The IUPAC name of neopentane is [AIEEE-2009]
 - (1) 2, 2-dimethylpropane
 - (2) 2-methylpropane
 - (3) 2, 2-dimethy/butane
 - (4) 2-methylbutane
- Arrange the carbanions, (CH₃)₃C̄, C̄Cl₃.
 (CH₃)₂C̄H, C₅H₅C̄H₂, in order of their decreasing stability [AJEEE-2009]
 - (1) (CH₂)₂CH > CCI₃> C₆H₅CH₂ > (CH₂)₃C
 - (2) CC13 > C8H8 CH2 > (CH3)2 CH > (CH3)5 C
 - (3) (CH3)3 C > (CH3)2 CH > C8H2 CH2 > CC13
 - (4) C₆H₅CH₂ > CCl₅ > (CH₅)₅C > (CH₅)₅CH
- The alkene that exhibits geometrical isomerism is [AJEEE-2009]
 - (1) 2 methyl propene
 - (2) 2 butene
 - (3) 2 methyl 2 butene
 - (4) Propene
- The correct order of increasing basicity of the given conjugate bases (R = CH₂) is [AIEEE-2010]
 - (1) RCOÖ<HC≡C<ÑH₂<R</p>
 - (2) RCOŌ < HC ≡ C̄ < R̄ < N̄H₅
 - (3) R<HC=C<RCOO<NH2
 - (4) RCOŌ<ÑH₅<HC≡Č<R
- 29.5 mg of an organic compound containing nitrogen was digested according to Kjeldahl's method and the evolved ammonia was absorbed in 20 mL of 0.1 M HCl solution. The excess of the acid required 15 mL of 0.1 M NaOH solution for

complete neutralisation. The percentage of nitrogen in the compound is [AIEEE-2010]

(1) 29.5

(2) 59.0

(3) 47.4

- (4) 23.7
- 6. The order of stability of the following carbocations

[JEE (Main)-2013]

- (1) 111 > 11 > 1
- (2) |1| > |1| > 1

- (3) 1 > 11 > 111
- (4) 111 > 1 > 11
- A gaseous hydrocarbon gives upon combustion 0.72 g of water and 3.08 g of CO₂. The empirical formula of the hydrocarbon is [JEE (Main)-2013]
 - (1) C₂H₄

(2) C₃H₄

(3) C₆H₅

- (4) C₂H_a
- For the estimation of nitrogen, 1.4 g of an organic compound was digested by Kjeldahl method and the evolved ammonia was absorbed in 60 mL of

M sulphuric acid. The unreacted acid required 20

mL of M sodium hydroxide for complete

neutralization. The percentage of nitrogen in the compound is [JEE (Main)-2014]

- (1) 6%
- (2) 10%
- (3) 3%
- (4) 5%
- In Carius method of estimation of halogens, 250 mg of an organic compound gave 141 mg of AgBr.
 The percentage of bromine in the compound is (At. mass Ag = 108; Br = 80) [JEE (Main)-2015]
 - (1) 24

(2) 36

(3) 48

(4) 60

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Arshad Iqbal

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