# **Acids and Bases**



Acid				Base		
Strength of Acid	Ka	Name	Formula	Formula	Name	Strengti of Base
trongest	Large	Perchloric acid	HCIO,	CIO,-	Perchlorate ion	Weakest
	3.2×10*	Hydrolodic acid		1-	Iodide	
	1.0×10*	Hydrobromic acid	HBr	Br-	Bromide	
	$1.3 \times 10^{\circ}$	Hydrochloric acid	HCI	CI-	Chloride	
	1.0×10°	Sulfuric acid	H,SO.	HSO	Hydrogen sulfate ion	
	2.4×10°	Nitric acid	HNO <sub>3</sub>	NO <sub>y</sub> -	Nitrate ion	
		Hydronium ion	H <sub>1</sub> O+	H <sub>2</sub> O	Water	
	5.4×10°	Oxelic acid	HO <sub>1</sub> C <sub>2</sub> O <sub>2</sub> H	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> -	Hydrogen oxalate ion	
	1.3×10-	Sulfurous acid	H <sub>3</sub> SO <sub>3</sub>	HSO <sub>3</sub> -	Hydrogen sulfite ion	
	1.0=10-	Hydrogen sulfate ion	HSO <sub>2</sub> -	50, 4	Sulfate ion	
	$7.4\times 10^{-3}$	Phosphoric acid	н, Ро.	H <sub>2</sub> PO <sub>2</sub> -	Dihydrogen phosphate ion	
	7.2×10*	Nitrous acid	HNO,	NO <sub>2</sub> -	Nitrite ion	
	6.6×10*	Hydrofluoric acid	HF	F-	Fluoride ion	
	1.8×10 <sup>-4</sup>	Methanoic acid	HCO <sub>2</sub> H	HCO <sub>2</sub> -	Methanoate ion	
	6.3×10°	Benzoic acid	С"Н,СООН	C,H,000-	Benzoate ion	
Medium	5.4×10°	Hydrogen oxalate ion	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> -	O <sub>2</sub> C <sub>2</sub> O <sub>2</sub> 2-	Oxalate ion	Medium
	1.8×10-4	Ethanoic acid	СН,СООН	CH,COO	Ethanoate (acetate) ion	
	$4.4 \times 10^{-7}$	Carbonic acid	н, со,	HCO <sub>2</sub> -	Hydrogen carbonate ion	
	$1.1 \times 10^{-3}$	Hydrosulfuric acid	H,5	HS-	Hydrogen sulfide ion	
	6.3×10*	Dihydrogen phosphate ion	H <sub>2</sub> PO <sub>4</sub> -	HPO. *-	Hydrogen phosphate ion	
	6.2×10*	Hydrogen sulfite ion	HS-	50,*	Sulfite ion	
	2.9×10 <sup>-4</sup>	Hypochlorous acid	HCIO	CIO-	Hypochlorite ion	
	$6.2 \times 10^{-10}$	Hydrocyanic acid	HCN	ON-	Cyanide ion	
	5.8×10++	Ammonium ion	NH. +	NH <sub>3</sub>	Ammonia	
	5.8×10 **	Boric acid	н,во,	H <sub>2</sub> BO <sup>2</sup>	Dihydrogen borate	
	$-4.7 \times 10^{-11}$	Hydrogen carbonate ion	HCO <sub>2</sub> +	COal-	Carbonate ion	
	$4.2 \times 10^{-10}$	Hydrogen phosphate ion	HPO, P	PO. F	Phosphate ion	
	$18\!\times\! 10^{-10}$	Dihydrogen borate ion	н,воз-	нво, н	Hydrogen borate ion	
	$1.3 \times 10^{-10}$	Hydrogen sulfide ion	HS-	5.6	Sulfide Ion	
	1.6×10-**	Hydrogen borate ion	нво, -	BO <sub>3</sub> *	Borate ion	
Weakest		water	H,O	OH-	Hydroxide	Stronges



# **Chemistry Acids And Bases Guide**

**Tao Wei** 

**Chemistry Acids And Bases Guide:** 

The Enigmatic Realm of Chemistry Acids And Bases Guide: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **Chemistry Acids And Bases Guide** a literary masterpiece penned with a renowned author, readers embark on a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book is core themes, assess its distinct writing style, and delve into its lasting effect on the hearts and minds of those that partake in its reading experience.

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