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- 2-1** (a) Stoichiometry is the description of the quantitative relationships among elements in a compound and among substances as they undergo chemical change.
(b) Composition stoichiometry describes the quantitative relationships among elements in compounds, e.g., in water, H_2O , there are 2 hydrogen atoms for every 1 atom of oxygen. Reaction stoichiometry describes the quantitative relationships among substances as they undergo chemical changes. (Reaction stoichiometry will be discussed in Chapter 3.)

The common ions for each formula unit is listed below:

(a) MgCl_2 contains Mg^{2+} and Cl^- ions

(b) $(\text{NH}_4)_2\text{CO}_3$ contains NH_4^+ and CO_3^{2-} ions

(c) $\text{Zn}(\text{NO}_3)_2$ contains Zn^{2+} and NO_3^- ions

2-5 Ethanol- $\text{CH}_3\text{CH}_2\text{OH}$



(space-filling; ball-and-stick)

Methanol- CH_3OH



(space-filling; ball-and-stick)

Both are composed of hydrogen, carbon, and oxygen. Both have an oxygen and hydrogen on the end. The ethanol molecule has an additional carbon and two hydrogens.

- 2-7** Organic compounds are those that contain carbon-to-carbon bonds, carbon-to-hydrogen bonds, or both. Organic formulas given in Table 2-1 include: acetic acid- CH_3COOH , methane- CH_4 , ethane- C_2H_6 , propane- C_3H_8 , butane- C_4H_{10} , pentane- C_5H_{12} , benzene- C_6H_6 , methanol- CH_3OH , ethanol- $\text{CH}_3\text{CH}_2\text{OH}$, acetone- CH_3COCH_3 , diethyl ether- $\text{CH}_3\text{CH}_2\text{COCH}_2\text{CH}_3$.

- 2-9** Compounds from Table 2-1 that contain only carbon and hydrogen and are not shown in Figure 1-5:

Compound	Ball and stick model	Compound	Ball and stick model
acetic acid- CH_3COOH		acetone- CH_3COCH_3	

Chemistry 9th Whitten Solution Manual

Wendy L. Keeney-Kennicutt



Chemistry 9th Whitten Solution Manual:

Student Solutions Manual for Whitten/Davis/Peck/Stanley's Chemistry, 9th Kenneth W. Whitten,Raymond E. Davis,Larry Peck,George G. Stanley,2009-04-20 Authored by Wendy L Keeney Kennicutt of Texas A M University this manual contains answers and solutions to all even numbered end of chapter exercises Solutions are divided by section for easy reference With this guide the author helps students achieve a deeper intuitive understanding of the material through constant reinforcement and practice **Subject Guide to Children's Books in Print 1997** Bowker Editorial Staff,R R Bowker Publishing,1996-09 *Subject Guide to Books in Print* ,2001 **Books in Print Supplement** ,1988 **Study Guide for Whitten/Davis/Peck/Stanley's Chemistry, 9th** Larry Peck,Raymond E. Davis,Kenneth W. Whitten,George G. Stanley,2009-05 By Raymond E Davis of the University of Texas Austin and James A Petrich of San Antonio College This study guide includes chapter summaries that highlight the main themes study goals with section references lists of important terms a preliminary test for each chapter that provides an average of 80 drill and concept questions and answers to the preliminary tests The Study Guide helps students organize the material and practice applying the concepts of the core text **Student Solutions Manual for Whitten, Davis, Peck, and Stanley's General Chemistry, 7th Ed** Wendy L. Keeney-Kennicutt,2004 **Government Reports Announcements & Index** ,1990 **Student Solutions Manual for Whitten, Davis, Peck, and Stanley's Chemistry** Wendy L. Keeney-Kennicutt,2010 General Chemistry With Solutions Manual Whitten,1996-09-01 **Study Guide for Whitten, Davis, Peck and Stanley's Chemistry, Tenth Edition** James Albert Petrich,Kenneth W. Whitten,Raymond E. Davis,Merlin Larry Peck,George G. Stanley,2014

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