

Chemical Reactions: Guided Notes

What is a chemical reaction?

1. A chemical reaction is a process that leads to the **transformation** of one set of chemical substances to another.
2. **Reactants** are the substances that *start* a chemical reaction.
3. **Products** are the substances that are *produced* in the reaction.

Evidence of a Chemical Reaction

Observing these clues can indicate that a chemical reaction has occurred:

1. **Change in Temperature:** Reactions can release heat (**exothermic**) or absorb it (**endothermic**).
2. **Formation of a Precipitate:** A **precipitate** is a solid that forms from a solution during a chemical reaction.
3. **Production of Gas:** Some reactions release gases, which you might see as **bubbles**.
4. **Change in Color:** A change in color often signals a new substance has formed.
5. **Change in Odor:** A change in **odor** can also indicate a new substance has formed.

Chemical Equations

1. A chemical equation is a symbolic representation of a chemical reaction.
2. **(s)**, **(l)**, **(g)**, and **(aq)** are symbols used to indicate the physical state of the substances involved in a reaction.
3. The **arrow** in a chemical equation separates the reactants from the products.

Balancing Chemical Equations

1. The **law of conservation of mass** states that matter cannot be created or destroyed in a chemical reaction.
2. To balance a chemical equation, you must have the same number of **atoms** of each element on both sides of the equation.
3. You can balance an equation by changing the **coefficients** in front of the chemical formulas.

Types of Chemical Reactions

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A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the Speedy Publishing text.

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