CHEMICAL QUANTITIES REVIEW QUESTIONS

- Calculate the following:
 - a) How many moles are present in 8.5 g of magnesium nitrate, Mg(NO₃)₂?
 - b) How many atoms are there in 8.5 g of Mg(NO1)2?
 - c) How many moles are there in 4.20 x 10²¹ molecules of methane?
 - d) What is the mass of 1.0 x 10²³ molecules of water?
- Titanium reacts with oxygen to produce titanium oxide, the pigment that gives paint and sunscreen a white colour.
 - a) If 15.6 g of titanium reacts with oxygen to form 20.0 g of titanium oxide, calculate the mass of oxygen that reacted. What chemistry law are you using?
 - b) Calculate the percent composition of titanium oxide.
 - b) How much titanium oxide can be produced from 500 g of titanium and unlimited oxygen? What law are you using to solve this?
- What is the percent composition of: a) sodium carbonate (Na₂CO₃)?
 b) aluminum oxide (Al₂O₁)?
- 4. What is the empirical formula of a compound that is:
 - a) 30.4% N, 69.6% O (by mass)
 - b) 43.6% P, 56.4% O (by mass)
- What is the molecular formula of a compound with molar mass 60.0 g and composed of 39.97% C, 6.73% H, and 53.30% O (by mass)?
- A certain compound has the following percent composition: C 57.1 %, H 4.8%, and 0 - 38.1 %. If the molar mass is 126 g, what is the molecular formula?
- When 25.0 g of a certain organic compound containing C, H and O is subjected to combustion analysis, 27.8 g of CO₂ and 19.9 g of H₂O are recovered. Determine the empirical formula of the compound.
- How many grams of lead (II) chloride will be produced when 6.7 g of lead (II) nitrate react with excess hydrochloric acid to form nitric acid and lead (II) chloride?
- A reaction involved in the production of iron from iron ore is:
 - Fe₂O₃(s) + CO(g) Fe(s) + CO₂(g)
 - a) How many kilograms of CO must react to produce 3.50 kg of Fe?
 - b) What mass of CO2 would be produced?
- a) What is the maximaum amount of iron (III) oxide than can be produced when 255 g of iron (II) sulfide are reacted with 71.4 g of oxygen gas in a closed container to produce iron (III) oxide and sulfur dioxide gas?
 - b) What mass of the excess reactant remains after the reaction?

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