

Chemical Basis of Life

Vocabulary

Polarity

Concentration

4 carbon compounds

monosaccharide

lipid

enzyme

Hydrogen bond

Solution

monomer

polysaccharide

fatty acid

catalyst

Adhesion

Solute

polymer

amino acid

nucleic acid

pH

Cohesion

Solvent

carbohydrate

protein

nucleotide

specific heat

Concepts to Know

WATER

Why so special and important?

Nearly 75% of our earth is covered in water.

Our bodies are 60-65% water. Our blood is 85% water.

None of the many reactions that occur in our bodies would occur without water.

Composition of water: H_2O

Properties of Water

- Water is **POLAR**.

The diagram shows that one side of the water molecule is **positively charged** and the other is **negatively charged**.

The molecule itself is neutral. These opposite charges make water a **polar molecule**.

- Water undergoes **ADHESION** and **COHESION** due to its polarity.

The **negative oxygen** of one water molecule is attracted to the **positive hydrogen** of another molecule forming a **hydrogen bond**. In other words water likes to stick to itself. Water sticking to water is called **cohesion**. Water's polarity allows it to attract to other surfaces too (like glass). Water sticking to something else is called **adhesion**.

- In the diagram to the right use dotted lines to draw in the bonds that form between water molecules.

Positives connect to negatives, up to four bonds per water molecule.

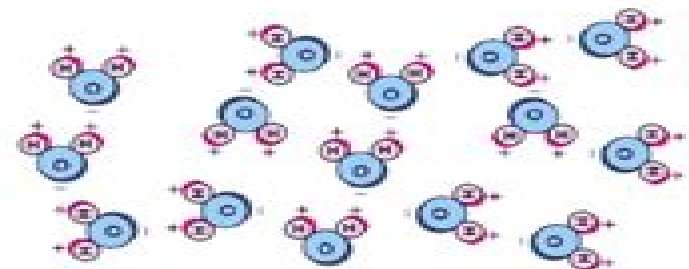
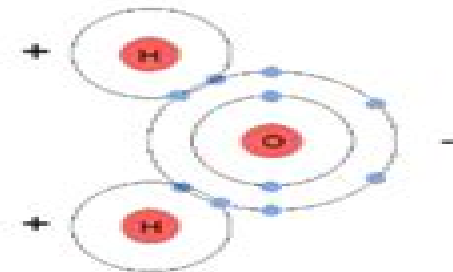
- What is the name of this type of bond?

Hydrogen or "H" bond

- What special property do the bonds give to water.

Several: cohesion first and foremost (water sticking to itself).

Creates surface tension on the water (water skipper can seemingly "walk" on water)



Chemical Basis Of Life Answer Key

JS Bruner



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1. Understanding the eBook Chemical Basis Of Life Answer Key
 - The Rise of Digital Reading Chemical Basis Of Life Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Basis Of Life Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Basis Of Life Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Basis Of Life Answer Key
 - Personalized Recommendations
 - Chemical Basis Of Life Answer Key User Reviews and Ratings
 - Chemical Basis Of Life Answer Key and Bestseller Lists
5. Accessing Chemical Basis Of Life Answer Key Free and Paid eBooks
 - Chemical Basis Of Life Answer Key Public Domain eBooks

- Chemical Basis Of Life Answer Key eBook Subscription Services
- Chemical Basis Of Life Answer Key Budget-Friendly Options
- 6. Navigating Chemical Basis Of Life Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Basis Of Life Answer Key Compatibility with Devices
 - Chemical Basis Of Life Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Basis Of Life Answer Key
 - Highlighting and Note-Taking Chemical Basis Of Life Answer Key
 - Interactive Elements Chemical Basis Of Life Answer Key
- 8. Staying Engaged with Chemical Basis Of Life Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Basis Of Life Answer Key
- 9. Balancing eBooks and Physical Books Chemical Basis Of Life Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Basis Of Life Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Basis Of Life Answer Key
 - Setting Reading Goals Chemical Basis Of Life Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Basis Of Life Answer Key
 - Fact-Checking eBook Content of Chemical Basis Of Life Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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