

## Review: Covalent Bonding

1. (Valence/Core) electrons are (shared/transferred) in the formation of an ionic bond.
2. (Valence/Core) electrons are (shared/transferred) in the formation of a covalent bond.
3. Which atom is placed in the center when drawing Lewis dot structures.
  - a. Fluorine or Carbon
  - b. Chlorine or Nitrogen
  - c. Nitrogen or Hydrogen
4. How did you determine which atom went in the center?

5. Complete the table.

| Lewis Dot Structure               | How many electrons are being shared? |
|-----------------------------------|--------------------------------------|
| $\text{:N}\equiv\text{N:}$        |                                      |
| $\text{H}-\ddot{\text{F}}:$       |                                      |
| $\ddot{\text{O}}=\ddot{\text{O}}$ |                                      |

6. Complete the table.

| Type of bond | Properties |
|--------------|------------|
| Ionic        |            |
| Covalent     |            |
| Metallic     |            |

7. Bonds are formed so that each atom achieves the octet rule. What does that mean?
8. The 1st step in drawing a Lewis dot structure is to determine the number of \_\_\_\_\_ of each atom?
9. What does VSEPR refer to?
10. Label the following diagrams as ionic, polar, or nonpolar.

# Covalent Bonding Core Teaching Resources Test Key

**L Towne**



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