Review: Covalent Bonding

- (Valence/Core) electrons are (shared/transferred) in the formation of an ionic bond.
- 2. (Valence/Core) electrons are (shared/transferred) in the formation of a covalent bond.
- 3. Which atom is placed in the center when drawing Lewis dot structures.
 - a. Fluorine or Carbon.
 - b. Chlorine or Nitrogen
 - c. Nitrogen or Hydrogen
- 4. How did you determine which atom went in the center?
- 5. Complete the table.

Lewis Dot Structure	How many electrons are being shared?
:N≡N:	
н— <u>Ё</u> :	
O = O	

6. Complete the table.

Type of bond	Properties
Ionic	
Covalent	
Metallic	

- 7. Bonds are formed so that each atom achieves the octet rule. What does that mean?
- 8. The 1st step in drawing a Lewis dot structure is to determine the number of _______ of each atom?
- 9. What does VSEPR refer to?
- 10. Label the following diagrams as ionic, polar, or nonpolar.

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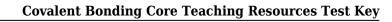
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