Standardized Test prep Answers

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1. C	4. 8	7. C	1.8 4.8	7. A							
E. C.	5. C	B. A	2. C 5. A	8.0							
1.0	6. C	9.8	3. C 6. 6	9.1							
io. Answers	many waters		10. 2 cups: d	linect proport	tion						
11. Observations C and G are chemical properties; the others are physical properties. 12. In a chemical change, one or more substances are converted into different substances. A physical change does not involve a change in the identity of the substance or substances present. 13. metals: shiny; good conductors of heat; good-conductors of electricity; malleable or ductile; most are solids at room temperature nonmetals: poor-conductors of heat; poor-conductors of electricity; many are gases at room temperature; those that are solids are brittle rather than malleable or ductile metalloids: properties intermediate between those of metals and nonmetals; less malleable than metals but			11. Statements A and B contain exact numbers. 12. The type of fertilizer is the variable being tested. Control factors are the types of radishes, the amount of water and the amount of sunshine. One control row should be planted under the same control factors but with no fertilizer. There are at least four things that could be used to determine the results: size, quantity, appearance and taste. Analysis might include bar graphs of each of these measurements for each of the five fertilizer types and the nofertilizer control row. 13. A unit must be defined in a way that does not depend on the circumstances of the measurement. Not every thumbnall is the same								
						not as brittle		most are semiconductors of electricity	stre.		7
							Chapte	r 3 Page 93	Chap 4 Pa	ge 129	
1. C	4. A 7. B		1. A	4. D	7. C						
2. D	5.8 8.0		2. 0	5.8	8. 0						
B. C.	6. A 9. C		3. A.	6. 8	9. C						
 Argon-40 has 22 neutrons (40 - 18 = 22), and potassium-40 has 21 neutrons (40 - 19 = 21). 6.17 g All cathode rays are the same, regardless of their source. Therefore, the particles responsible for the cathode rays must be present in all atoms. The particles are electrons. When the average atomic mass is calculated, it is 10.811. Because the atomic mass is the same as the atomic mass of boron, mythium was not a new element. 			10. The color will be orange. Converting energy into frequency gives 4.8 × 10 ¹⁴ , which corresponds to the frequency of orange light. 11. 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1								
Chapter 5 Page 1			Chapter 6	Page 215							
1. C			3. A 2	2. C	7. A						
2. A.			3.8 4	L D	8.6						
1. A.			5. A 6	5. D	9. C						
6. A.			10. Hybridiza	ation explain	s how the orbitals of an atom become						
s. C					m forms covalent bonds.						
5. A.					the because shifting of the layers of ions result						
7. D					that cause the layers to part completely.						
1.0					stronger intermolecular forces even though i						
9. C			is non-polar, Boiling point	because its I is directly or	boilling point is higher than that of acetic acid orrelated to strength of intermolecular forces						
 5: 2251 ki/mol; Cl: 2297 ki/mol; Ar: 2666 ki/mol; K: 3051 ki/mol; Ca: 1145 ki/mol; Sc: 1235 ki/mol; Ti: 1310 ki/mol. For the second 					lecular forces, the more energy needed to lar forces, and therefore the higher the boilir						

hongsh.

point. Naphthalene is so large that its dispersion forces are greater than

 There is a repulsive region at small bond distances, a minimum well at an intermediate bond distance, and an asymptotic region at large

the sum of the dispersion forces and hydrogen bonding in acetic acid.

bond distances. The regulaive region is due to the regulation between

other. The asymptotic region is where the atoms are far apart and therefore do not interact with each other. The potential energy decreases as the atoms come closer together and are attracted to each other. The minimum well is the region where the attractive forces are balanced by the repulsive forces. This distance is the equilibrium bond

the two positive nuclei at distances where the nuclei are close to each

Ca: 1145 kl/mol; Sc: 1235 kl/mol; Ti: 1310 kl/mol. For the second ionization, the general trend is for increasing IE2 across the period in Groups 2–18 with Group 1 having the highest IE2. IE2 decreases going down a group.

 Group 16 most commonly forms 2- ions, because these elements require only two more electrons to fill their shell (obtain a noble-gas configuration).

Chemistry Standardized Test Prep Answers

M Walker

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