



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time having reached Equilibrium. Chemical reactions are represented by the reaction: $\rightleftharpoons$

What is Equilibrium? Chemical Equilibrium is a state where the concentrations of the chemical species remain the same with time having reached Equilibrium. Chemical reactions are represented by the reaction: $\rightleftharpoons$ Equilibrium Constant K_c Equilibrium Constant K_p	Equilibrium Constant Equilibrium Constant K_c Equilibrium Constant K_p	
Equilibrium Constant K_c Equilibrium Constant K_c is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. Equilibrium Constant K_c	Equilibrium Constant K_p Equilibrium Constant K_p is defined as the ratio of the partial pressures of the products to the partial pressures of the reactants, each raised to the power of its stoichiometric coefficient. Equilibrium Constant K_p	
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Le Chatelier's Principle: the effect of changing conditions on equilibrium		
Concentration Increasing the concentration of a reactant or product will shift the equilibrium position to the right or left, respectively.	Pressure Increasing the pressure will shift the equilibrium position to the side with the fewest moles of gas.	Temperature Increasing the temperature will shift the equilibrium position to the side that is endothermic.
Concentration Increasing the concentration of a reactant or product will shift the equilibrium position to the right or left, respectively.	Pressure Increasing the pressure will shift the equilibrium position to the side with the fewest moles of gas.	Temperature Increasing the temperature will shift the equilibrium position to the side that is endothermic.
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