

PROBLEMS WITH CALCULATIONS

1. Nitrogen gas and hydrogen gas combine according to the balanced equation below:



The diagram below shows an initial mixture of nitrogen gas (blue molecules) and hydrogen gas (red spheres) which react to form ammonia gas (red spheres). Before the synthesis of the final product molecule, explain what happens.



Initial reactant mixture



Final mixture mixture

Explanation:

There are 4 N atoms in the initial mixture, coming from 2 molecules of N_2 . And 12 H atoms, coming from 6 molecules of H_2 . And 12 H atoms, coming from 2 molecules of NH_3 . And 4 N atoms, coming from 2 molecules of NH_3 . And 6 H atoms, coming from 3 molecules of H_2 . And 4 N atoms, coming from 2 molecules of NH_3 . And 6 H atoms, coming from 3 molecules of H_2 . And 4 N atoms, coming from 2 molecules of NH_3 . And 6 H atoms, coming from 3 molecules of H_2 .

2. Use the same equation as question 1 to answer the following questions.

- a. If 28.0 g of N_2 and 4.0 g of H_2 are reacted, what mass of NH_3 can be produced?

$$\begin{aligned} \text{N}_2 + 3\text{H}_2 &\rightarrow 2\text{NH}_3 \\ 28.0 \text{ g N}_2 &\times \frac{1 \text{ mol N}_2}{28.0 \text{ g N}_2} \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \times \frac{17.0 \text{ g NH}_3}{1 \text{ mol NH}_3} = 34.0 \text{ g NH}_3 \\ 4.0 \text{ g H}_2 &\times \frac{1 \text{ mol H}_2}{2.0 \text{ g H}_2} \times \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \times \frac{17.0 \text{ g NH}_3}{1 \text{ mol NH}_3} = 22.7 \text{ g NH}_3 \end{aligned}$$

- b. What is the limiting reactant? Explain your work.

Answer: H_2 is the limiting reactant.

Answer: H_2

Answer: H_2

- c. What is the percent yield if 34.0 g of ammonia is produced in this reaction?

$$\frac{34.0 \text{ g}}{34.0 \text{ g}} \times 100\%$$

Answer: 100%

Ccsf Chem 40 Lab Manual

Marquis Who's Who



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