

Chapter 21

Organic Chemistry: Some Basic Principles and Techniques

1. The IUPAC name of neopentane is [AIEEE-2009]

(1) 2, 2-dimethylpropane
(2) 2-methylpropane
(3) 2, 2-dimethylbutane
(4) 2-methylbutane

2. Arrange the carbanions, $(\text{CH}_3)_3\text{C}^-$, CCl_3^- , $(\text{CH}_3)_2\text{CH}^-$, $\text{C}_6\text{H}_5\text{CH}_2^-$, in order of their decreasing stability [AIEEE-2009]

(1) $(\text{CH}_3)_2\text{CH}^- > \text{CCl}_3^- > \text{C}_6\text{H}_5\text{CH}_2^- > (\text{CH}_3)_3\text{C}^-$
(2) $\text{CCl}_3^- > \text{C}_6\text{H}_5\text{CH}_2^- > (\text{CH}_3)_2\text{CH}^- > (\text{CH}_3)_3\text{C}^-$
(3) $(\text{CH}_3)_3\text{C}^- > (\text{CH}_3)_2\text{CH}^- > \text{C}_6\text{H}_5\text{CH}_2^- > \text{CCl}_3^-$
(4) $\text{C}_6\text{H}_5\text{CH}_2^- > \text{CCl}_3^- > (\text{CH}_3)_3\text{C}^- > (\text{CH}_3)_2\text{CH}^-$

3. The alkene that exhibits geometrical isomerism is [AIEEE-2009]

(1) 2-methyl propene
(2) 2-butene
(3) 2-methyl-2-butene
(4) Propene

4. The correct order of increasing basicity of the given conjugate bases ($\text{R} = \text{CH}_3$) is [AIEEE-2010]

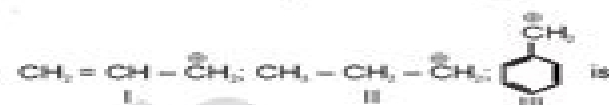
(1) $\text{RCOO}^- < \text{HC}=\text{C}^- < \text{NH}_2^- < \text{R}^-$
(2) $\text{RCOO}^- < \text{HC}=\text{C}^- < \text{R}^- < \text{NH}_2^-$
(3) $\text{R}^- < \text{HC}=\text{C}^- < \text{RCOO}^- < \text{NH}_2^-$
(4) $\text{RCOO}^- < \text{NH}_2^- < \text{HC}=\text{C}^- < \text{R}^-$

5. 29.5 mg of an organic compound containing nitrogen was digested according to Kjeldahl's method and the evolved ammonia was absorbed in 20 mL of 0.1 M HCl solution. The excess of the acid required 15 mL of 0.1 M NaOH solution for

complete neutralisation. The percentage of nitrogen in the compound is [AIEEE-2010]

(1) 29.5 (2) 59.0
(3) 47.4 (4) 23.7

6. The order of stability of the following carbocations



[JEE (Main)-2013]

(1) $\text{III} > \text{II} > \text{I}$ (2) $\text{II} > \text{III} > \text{I}$
(3) $\text{I} > \text{II} > \text{III}$ (4) $\text{III} > \text{I} > \text{II}$

7. A gaseous hydrocarbon gives upon combustion 0.72 g of water and 3.08 g of CO_2 . The empirical formula of the hydrocarbon is [JEE (Main)-2013]

(1) C_2H_6 (2) C_2H_4
(3) C_6H_6 (4) C_7H_8

8. For the estimation of nitrogen, 1.4 g of an organic compound was digested by Kjeldahl method and the evolved ammonia was absorbed in 60 mL of

$\frac{\text{M}}{10}$ sulphuric acid. The unreacted acid required 20

mL of $\frac{\text{M}}{10}$ sodium hydroxide for complete neutralization. The percentage of nitrogen in the compound is [JEE (Main)-2014]

(1) 6%
(2) 10%
(3) 3%
(4) 5%

9. In Carius method of estimation of halogens, 250 mg of an organic compound gave 141 mg of AgBr. The percentage of bromine in the compound is (At. mass Ag = 108; Br = 80) [JEE (Main)-2015]

(1) 24 (2) 36
(3) 48 (4) 60

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