

chemistry stoichiometry questions and answers

University of Binawan

1. Complete combustion of ammonia gas requires oxygen gas with equality reaction :



With Assuming that the oxygen level in the air is 20%, then the volume of air needed to burn 4 liters of ammonia gas is...

- A. 5 liters
- B. 7 liters
- C. 10 liters
- D. 35 liters

Answer: D

Discussion

Equality reaction equivalent :



Volume gas oxygen = $(7/4) \times 4$ liters = 7 liters

Volume oxygen = 20% volume air

7 liters = 0.2 volume of air

volume udara = $7/0.2 = 35$ litres

2. Every 1 liter of nitrogen gas is appropriate finished react with 2.5 liters of oxygen gas to form 1 liter of nitrogen oxide gas . If the volume is measured at the same temperature and pressure , then formula molecule the nitrogen oxide is

- A. N_2O
- B. NO
- C. N_2O_3 _ _ _
- D. N_2O_5 _

Answer: D

Discussion

Equality reaction :



The number of N atoms before reaction = Number of N atoms after reaction

$$4 = 2x$$

$$x = 2$$

Number of O atoms before reaction = Number of O atoms after reaction

$$10 = 2y$$

$$y = 5$$

So, formula molecule The nitrogen oxide is N_2O_5

3. Propane gas (C_3H_8) is burned perfect with reaction :



For burn perfect every liters of propane gas, oxygen gas is required as much

- A. 1.0 liters
- B. 3.0 liters

Chemistry Stoichiometry Guide Answers

Ian Guch



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