

Chemical Kinetics Lab Report

Title

Using Chemical Kinetics to Analyze Concentration and Reaction Rates

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Introduction

Chemical kinetics introduces the concept of understanding the rates of chemical reactions and, for the sake of this experiment, how the speed of molecular reconstruction is affected by an increase or decrease of a reactant's concentration. The rate of reaction quantitatively measures the change in the amount of reactant or product per unit of time (Flowers et al., 2015). To test this concept, this experiment analyzed the reaction between acetone (CH_3COCH_3) and iodine (I_2), represented by the following equation:

$CH_3COCH_3(aq) + I_2(aq) \rightarrow CH_3COCH_2I(aq) + H^+(aq) + I^-(aq)$. To enhance the significance of each chemical, the concentrations between acetone and iodine were altered amongst three trials and timed accordingly. The reaction was determined by the time the chemicals turned from yellow to colorless, which occurs due to the consumption of iodine as the reaction proceeds. By increasing an individual reactant and keeping the remaining chemicals constant during respected trials, it could be determined which concentration of a specific reactant held more of an effect on the reaction rate to find the relationship between reactants and reaction rates. Because acetone was known to have a greater surface area and evaporation rate compared to that of iodine, an initial claim was made that the fastest reaction rate would occur in the trial with an increase in the concentration of acetone.

Experimental Methods

Deionized water was introduced into the experiment with the knowledge that the liquid itself holds no effect on any chemical reaction. The total volume of each solution was to be measured out to 50mL; the addition of deionized water allowed for this experiment to avoid any inaccuracies. An additional liquid, hydrochloric acid, remained at a constant 10mL throughout the entirety of the experiment as it served as

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David Garvin



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Scientific and Technical Aerospace Reports ,1995 **SCR.** ,1968 **Reaction Rates for High-temperature Air with Carbon and Sodium Impurities** Mina L. Carnicom,1968 The values used by a number of investigators for the rate constants of high temperature greater than or equal to 1000 K homogeneous gaseous reactions involving species of the elements nitrogen oxygen carbon and sodium have been compiled and are presented in tabular form Included are reactions involving neutral species charged species free electrons some species in excited electronic or vibrational states and radiative processes U.S. Government Research Reports ,1964 *Subject Index to Unclassified ASTIA Documents* Defense Documentation Center (U.S.),1960 **Ninth Symposium (International) on Combustion** Sam Stuart,2013-09-11 Ninth Symposium International on Combustion covers the proceedings of the Ninth Symposium International on Combustion held at Cornell University in Ithaca New York on August 27 to September 1 1962 under the auspices of the Combustion Institute The book focuses on the processes and reactions involved in combustion The selection first offers information on flame strength of propane oxygen flames at low pressures in turbulent flow and mixing and flow in ducted turbulent jets Topics include radial profile of the jetting velocity radial growth of the jet and mixing zones of a ducted jet The text then elaborates on turbulent flame studies in two dimensional open burners turbulent mass transfer and rates of combustion in confined turbulent flames and flame stabilization in a boundary layer The publication examines the theoretical study of properties of laminar steady state flames as a function of properties of their chemical components and spectra of alkali metal organic halide flames The text then takes a look at the thermal radiation theory for plane flame propagation in coal dust clouds flame characteristics of the diborane hydrazine system and studies of the combustion of dimethyl hydrazine and related compounds The selection is a dependable reference for readers interested in the processes and reactions involved in combustion **A Compendium of Evaluated and Estimated Rate Coefficients** David Garvin,1968 Charts of Equilibrium Real-gas Stagnation Point Conditions on a Sphere for Altitudes to 200,000 Feet and Velocities to 15,000 Feet Per Second P. D. Fisher,1964 This report presents in graphical form the stagnation point pressure temperature equilibrium surface temperature and heating rate for a sphere Equilibrium real gas effects and surface radiation are considered The charts cover a range of velocities from 0 to 15 000 feet per second and altitudes from sea level to 200 000 feet Corrections for sphere diameter and emissivity are included Author *Government-wide Index to Federal Research & Development Reports* ,1965 **Combustion & Emission Formation Process in Diesel Engines** ,2003

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