

➔ CHEMICAL EQUILIBRIUM

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Q.1: Justify that chemical equilibrium is dynamic in nature?

Ans: In reversible chemical reactions the molecules of reactions collide and convert into products. At the same time the molecules of the products are converted into reactants. When two opposing forces maintain the equal, rates then equilibrium is there and that is dynamic equilibrium in nature.

Q.2: When does four types of chemical equilibrium constants for a reaction become equal?

Ans: When we have such a reversible reaction, in which the total number of moles of reactants and products are equal then $K_p = K_c = K_n$. In other words $\Delta n = 0$, where Δn is the difference of number of moles of reactants and products.

Q.3: The change of temperature disturbs the equilibrium position and the equilibrium constant of reaction? Justify?

Ans: All the reversible reactions are disturbed by changing their equilibrium position and equilibrium constant by disturbing the temperature. Actually change of temperature changes the energy contents of reactants and products.

Q.4: Why is solubility of glucose in water is increased by increasing the temperature?

Ans: The solubility of glucose in water is an endothermic process. Increase of temperature pushed the system to that side where heat is absorbed, and that is the side where solubility increases.

Q.5: What is the nature of solution having $\text{pH} = 7.5$?

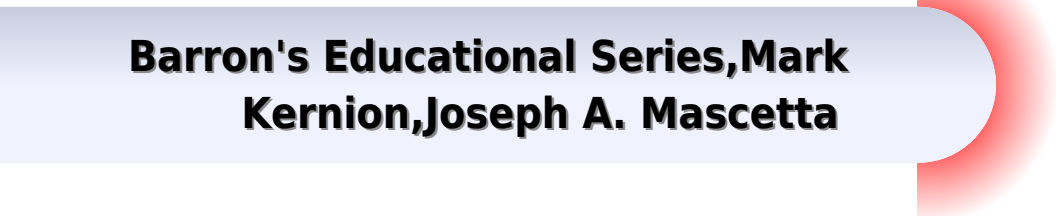
Ans: The solution with $\text{pH} = 7$ is neutral and having values more than 7 have greater concentration of OH^- and less concentration of H^+ , so the solution is basic.

Q.6: What is buffer capacity?

Ans: Buffer capacity is a ability of a buffer to resist the change of pH , when a few drops of an acid and base is added from outside. The best buffer is obtained. (i) when pH of the buffer is equal to pK_a of the acid and the concentration of the salt and acid are equal. (ii) The OH^- of the buffer is equal to pK_b or the base.

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**Barron's Educational Series, Mark
Kernion, Joseph A. Mascetta**



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