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- 2-1** (a) Stoichiometry is the description of the quantitative relationships among elements in a compound and among substances as they undergo chemical change.
(b) Composition stoichiometry describes the quantitative relationships among elements in compounds, e.g., in water, H_2O , there are 2 hydrogen atoms for every 1 atom of oxygen. Reaction stoichiometry describes the quantitative relationships among substances as they undergo chemical changes. (Reaction stoichiometry will be discussed in Chapter 3.)

The common ions for each formula unit is listed below:

(a) MgCl_2 contains Mg^{2+} and Cl^- ions

(b) $(\text{NH}_4)_2\text{CO}_3$ contains NH_4^+ and CO_3^{2-} ions

(c) $\text{Zn}(\text{NO}_3)_2$ contains Zn^{2+} and NO_3^- ions

2-5 Ethanol- $\text{CH}_3\text{CH}_2\text{OH}$



(space-filling; ball-and-stick)

Methanol- CH_3OH



(space-filling; ball-and-stick)

Both are composed of hydrogen, carbon, and oxygen. Both have an oxygen and hydrogen on the end. The ethanol molecule has an additional carbon and two hydrogens.

- 2-7** Organic compounds are those that contain carbon-to-carbon bonds, carbon-to-hydrogen bonds, or both. Organic formulas given in Table 2-1 include: acetic acid- CH_3COOH , methane- CH_4 , ethane- C_2H_6 , propane- C_3H_8 , butane- C_4H_{10} , pentane- C_5H_{12} , benzene- C_6H_6 , methanol- CH_3OH , ethanol- $\text{CH}_3\text{CH}_2\text{OH}$, acetone- CH_3COCH_3 , diethyl ether- $\text{CH}_3\text{CH}_2\text{COCH}_2\text{CH}_3$.

- 2-9** Compounds from Table 2-1 that contain only carbon and hydrogen and are not shown in Figure 1-5:

Compound	Ball and stick model	Compound	Ball and stick model
acetic acid- CH_3COOH		acetone- CH_3COCH_3	

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S Marginson



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