

Diagnostische toets

H9 Redoxchemie

Deze toets bestaat uit 13 vragen.

Bij deze toets kun je maximaal 28 punten behalen.

Voor elk vraagnummer zijn de te behalen punten aangegeven.

Kerncentrale

Uraan(IV)oxide, $\text{UO}_2(\text{s})$, wordt gewonnen uit uraanerts voor gebruik in kerncentrales.

Uit uraanerts wordt eerst een ander uraanoxide verkregen: $\text{U}_3\text{O}_8(\text{s})$. In dit uraanoxide zijn uraanionen met verschillende elektrovalenties aanwezig. $\text{U}_3\text{O}_8(\text{s})$ wordt met behulp van een redoxreactie omgezet in uraan(IV)oxide, $\text{UO}_2(\text{s})$.

- 3p **1** Stel zelf de halfreactie op van de omzetting van $\text{U}_3\text{O}_8(\text{s})$ in $\text{UO}_2(\text{s})$. In de halfreactie komen naast de genoemde stoffen ook H_2O en H^+ voor.
- 2p **2** Leg uit of voor de volledige omzetting van uraanoxide, $\text{U}_3\text{O}_8(\text{s})$, in uraan(IV)oxide, $\text{UO}_2(\text{s})$, een oxidator of reductor nodig is.

Zink

Zink is geen edelmetaal, maar toch hoeven zinken voorwerpen niet te worden beschermd tegen corrosie.

- 2p **3** Geef hiervoor een verklaring.

Zink kan op verschillende manieren als bescherm laag worden aangebracht. Een voorwerp kan thermisch of elektrolytisch worden verzinkt. Het voordeel van elektrolytisch verzinken is dat het een zeer dun en glad laagje kan aanbrengen. Bij een bedrijf kan men staal voorzien van een laagje zink van slechts enkele micrometers dik. Hiervoor wordt het stalen voorwerp twaalf uur lang elektrolytisch verzinkt in een zinksulfaatoplossing met een stroomsterkte van 2,5 A.

- 4p **4** Bereken hoeveel gram zink hierbij wordt verbruikt.

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A. D. de Groot, R. F. van Naerssen



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