# CHEMICAL EQUILIBRIUM

$$K_c = \frac{[B]}{[A]} = \text{mol } L^{-1}$$

$$\succ$$
  $K_p = \frac{P_p}{P_A} = Partial Pressure$ 

$$K_n = \frac{X_n}{X_n} = \text{mole fraction}$$

$$ightharpoonup K_p = K_c(RT)^{torg}$$
 where  $\Delta n_g = n_p - n_g$ 

#### Predicting the extent of reaction :

- K<sub>C</sub> > 10<sup>3</sup> [Forward reaction is favoured.]
- K<sub>c</sub> < 10<sup>-3</sup> [Reverse reaction is favoured.]
- 10<sup>-3</sup> < K<sub>C</sub> < 10<sup>3</sup> [Both reactants and products are present in equilibrium]

#### Free Energy Charge (△G)

- a) If ΔG = 0 then reversible reaction would be in equilibrium, K<sub>c</sub> = 1
- b) If AG = (+) ve then equilibrium will be displaced in backward direction; K<sub>c</sub> < 1</li>
- c) If AG = (-) we then equalibrium will shift in forward direction;

$$K_c > 1$$

(a) K<sub>c</sub> unit → (moles/lit)<sup>to</sup>

AG or a reaction under any set of conditions is related to its value under standard conditions, i.e. AG° by the equation

$$\Delta G = \Delta G^{\circ} + 2.303 \text{ RT log Q}$$

 Under equlibrium conditions, for same number of moles of reactants and products

$$Q = K_p = K_c = K$$
 and  $\Delta G = 0$ 

$$\Delta G^{\circ} = -2.303 \text{ RT log K}$$

 We have replaced K<sub>p</sub> by K called thermodyannic equilibrium constant

## Significance of AG°

Significance of AG° can be explained from the following points

(i) If ∆G° < 0, log K > 0, ⇒ K > 1

Hence, reaction is spontaneas in forward direction.

(ii) Hence ∆G° < 0, log K > 0, ⇒ K < 1</p>

Hence, reaction is non-spontaneous or a reaction proceeds in the forward direction to such a small proceeds in the forward direction to such a small extent that a very small amount of the product is formed.

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Marcia L. Gillette,Ira Blei,Wendy Gloffke,George Odian

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