

DIAGNOSTIC TEST

DIAGNOSTIC TEST

Time: 60 minutes

Part A

Note: Unless otherwise stated, for all statements involving chemical equations and/or solutions, assume that the system is in pure water.

Directions: Each of the following sets of lettered choices refers to the numbered formulas or statements immediately below it. For each numbered item, choose the one lettered choice that fits it best. Then fill in the corresponding oval on the answer sheet. Each choice in a set may be used once, more than once, or not at all.

Questions 1–4

- (A) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$
- (B) $\text{CH}_3\text{CH}(\text{Br})\text{CH}(\text{Br})\text{CH}_3$
- (C) $\text{CH}_3\text{CH}_2\text{CH}(\text{Br})\text{CH}_3$
- (D) $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$
- (E) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$



(3)

4. Which compound would most likely turn litmus paper to a red color?

Questions 5–8

- (A) Heisenberg Uncertainty Principle
- (B) Pauli Exclusion Principle
- (C) Schrödinger Wave Equation
- (D) Hund's rule
- (E) Bohr model of the hydrogen atom

- 5. No two electrons can have the same quantum number because they must have opposite spins.
- 6. We cannot know the exact location of an electron in space.
- 7. The electrons will occupy an orbital singly, with parallel spins, before pairing up.

8. The energy changes that an electron may undergo are quantized.

Questions 9–12

- (A) H_2
- (B) CO_2
- (C) H_2O
- (D) NaCl
- (E) CH_3CH_3

- 9. Contains just one sigma bond.
- 10. Has a bond formed from the transfer of electrons.
- 11. Has an atom that is sp hybridized.
- 12. Is a polar molecule.

Questions 13–16

- (A) F
- (B) Li
- (C) Fe
- (D) He
- (E) Si

- 13. Shows both the properties of both metals and non-metals.
- 14. Has the greatest ionization energy.
- 15. Has the greatest electronegativity.
- 16. Has colored salts that will produce colored aqueous solutions.

Chemistry Diagnostic Test Topic 3 Answers

Samuel Ralph Powers



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Table of Contents Chemistry Diagnostic Test Topic 3 Answers

1. Understanding the eBook Chemistry Diagnostic Test Topic 3 Answers
 - The Rise of Digital Reading Chemistry Diagnostic Test Topic 3 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Diagnostic Test Topic 3 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Diagnostic Test Topic 3 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Diagnostic Test Topic 3 Answers
 - Personalized Recommendations
 - Chemistry Diagnostic Test Topic 3 Answers User Reviews and Ratings
 - Chemistry Diagnostic Test Topic 3 Answers and Bestseller Lists

5. Accessing Chemistry Diagnostic Test Topic 3 Answers Free and Paid eBooks
 - Chemistry Diagnostic Test Topic 3 Answers Public Domain eBooks
 - Chemistry Diagnostic Test Topic 3 Answers eBook Subscription Services
 - Chemistry Diagnostic Test Topic 3 Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Chemistry Diagnostic Test Topic 3 Answers Compatibility with Devices
 - Chemistry Diagnostic Test Topic 3 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Diagnostic Test Topic 3 Answers
 - Highlighting and Note-Taking Chemistry Diagnostic Test Topic 3 Answers
 - Interactive Elements Chemistry Diagnostic Test Topic 3 Answers
8. Staying Engaged with Chemistry Diagnostic Test Topic 3 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Diagnostic Test Topic 3 Answers
9. Balancing eBooks and Physical Books Chemistry Diagnostic Test Topic 3 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Diagnostic Test Topic 3 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Diagnostic Test Topic 3 Answers
 - Setting Reading Goals Chemistry Diagnostic Test Topic 3 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Diagnostic Test Topic 3 Answers
 - Fact-Checking eBook Content of Chemistry Diagnostic Test Topic 3 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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