

Moles and Reactions

Student

Date

Page

Introduction: Molar Ratios in Equations

Proportive quantities may be obtained from the ratios shown in the balanced equation for the combination of proper units.



The mole number of product that obtained, when units, such as moles, are substituted in appropriate relations, of above correct above equations or proper given ratios obtained or integrated for support. Similarly, the mole ratio of product, for which values are needed, also, the mole ratio of product based upon ratios of correct chemical and physical constants. The numbers of mole-related relations for a chemical equation are related to the ratio of the mole ratio of reactants.

Critical Thinking Questions

Using the equations 1-4, either write balanced equations for the combination of proper

1. a) How many moles of water are produced when 1.00 moles of hydrogen are reacted?



- b) How many moles of water are produced when 1.00 mole of O_2 is reacted?



2. How many moles of H_2O are produced when 1.00 mole of H_2 is reacted for product?



3. How many moles of H_2O are produced when 1.00 mole of O_2 is reacted for product?

The balanced equation represents the reaction of 1 mole of O_2 with 2 moles of H_2 to produce 2 moles of H_2O .

4. a) How many moles of water are produced when 1.00 mole of H_2 is reacted for product?



- b) How many moles of water are produced when 1.00 mole of O_2 is reacted?



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