

Standard
state
of
formation
of
product
is
1 mol

1. If thermodynamic measurements are made under the following conditions:
 1. All reactants and products are in their standard states.
 2. The pressure is 1 bar.
 3. The temperature is 298.15 K.
 4. The reaction is written as a balanced chemical equation.

- ① reactants are elements in most stable form
- ② substance is a product
- ③ you obtain exactly one mole of product

2. Standard molar enthalpy of formation calculation follows the following equation:
 $\Delta H_{\text{rxn}}^\circ = \sum \Delta H_{\text{f, products}}^\circ - \sum \Delta H_{\text{f, reactants}}^\circ$

Compound	ΔH _f (kJ/mol)
CO(g)	-110.5
H ₂ (g)	0
H ₂ O(g)	-241.8

$$\Delta H_{\text{rxn}}^\circ = \Delta H_{\text{f, products}}^\circ - \Delta H_{\text{f, reactants}}^\circ$$

For the reaction: $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$

$$\Delta H_{\text{rxn}}^\circ = 2 \Delta H_{\text{f, H}_2}^\circ + \Delta H_{\text{f, CO}}^\circ - (\Delta H_{\text{f, CH}_3\text{OH}}^\circ + 2 \Delta H_{\text{f, H}_2}^\circ)$$

$$\begin{aligned} \Delta H_{\text{rxn}}^\circ &= 2(-110.5) + (-110.5) - (-200.7) - 2(0) \\ &= -221.0 - 110.5 + 200.7 + 0 \\ &= -130.8 \text{ kJ} \end{aligned}$$

Chemistry Thermochemical Equations Study Guide Answers

DJ Losen



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