

balancing Equations

Name: _____

Date: _____

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Introduction/Questions

- Write the equation for the reaction of sodium and chlorine (chlorine gas) to form sodium chloride.



- Write the equation for the reaction of sodium nitrate and sodium chloride to produce sodium chloride and sodium nitrate.



Information: Subscripts and Coefficients

It is important to understand that when you have subscripts in a compound, the number of Cl₂ molecules is the subscript and the fact that there are two chlorine atoms bonded to one carbon.

A coefficient tells you how many molecules of a compound there are, that is a different way. For example, for the equation: "3 NaCl" the three is the coefficient. The three tells us that there are three molecules of sodium present. In this equation "1 NaCl" there are a total of 1 sodium atoms and 1 chlorine atom.

Balance Chemical Equations

- Balance the gas phase reaction between H₂ and O₂. 3/1

- Balance the gas phase reaction for each of the following:



Information: How To Balance Equations

Consider the reaction of sodium and chlorine to produce sodium chloride (halogenation):



Remember that when chlorine is the reactant it is always written as Cl₂. On the product side of the equation, there must always be a total of two atoms of chlorine, but on the product side there is only one atom of chlorine. Atomic masses are also important. In order to balance equations for gases, we must account for balancing the equation. This can be done, but, by adding a "2" to the product side:



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The Directory of Management Consultants, 2003 Kennedy Information (Firm), 2002

Chemquest 37 Answer Key Book Review: Unveiling the Power of Words

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