

CHAPTER 13

13.1 States of matter review

Gases: contain particles that are far apart, in random motion, and independent of one another.

- Little contact
- No attractive forces
- Random motion

Solids: contain particles very close in space and maintain a rigid shape. Significant attractive forces exist between particles.

- Close contact
- Strong attractive forces
- Rigid shape

Liquids: intermediate between gases and solids.

Contain particles close to one another but have fluidity (can assume the shape of a container).

Significant attractive forces exist between particles in a liquid.

- Close contact
- Some attractive forces
- Fluid shape

13.2 Properties of liquids

Evaporation

Evaporation or Vaporization:

Escape of molecules from the liquid to the gas phase.

Liquid Vapor

Kinetic Energy: Energy associated with movement.

1. Molecules in the liquid state have different kinetic energies (KE's).
2. Those with higher KE's can overcome attractive forces between particles and escape to the gas phase.
3. We can increase the kinetic energy (or movement of the molecules in liquid state by adding heat (boiling).

Sublimation

Phase change from the solid to gas phase that bypasses the liquid state.

Solid vapor

Examples:

Fusion

Molecules can go from the liquid to solid phase.

This process is called "fusion"

Ex. Liquid solid

Under certain circumstances, gas can transform directly into a solid. This process is called "**deposition**".

Example: Water vapor to ice - Water vapor transforms directly into ice without becoming a liquid, a process that often occurs on windows during the winter months.

Condensation

Molecules from the liquid phase can escape to the vapor phase through evaporation.

Chapter 13 Study Guide Gases Answers

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Table of Contents Chapter 13 Study Guide Gases Answers

1. Understanding the eBook Chapter 13 Study Guide Gases Answers
 - The Rise of Digital Reading Chapter 13 Study Guide Gases Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 13 Study Guide Gases Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 13 Study Guide Gases Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 13 Study Guide Gases Answers
 - Personalized Recommendations
 - Chapter 13 Study Guide Gases Answers User Reviews and Ratings
 - Chapter 13 Study Guide Gases Answers and Bestseller Lists
5. Accessing Chapter 13 Study Guide Gases Answers Free and Paid eBooks
 - Chapter 13 Study Guide Gases Answers Public Domain eBooks
 - Chapter 13 Study Guide Gases Answers eBook Subscription Services
 - Chapter 13 Study Guide Gases Answers Budget-Friendly Options
6. Navigating Chapter 13 Study Guide Gases Answers eBook Formats

- ePub, PDF, MOBI, and More
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- Chapter 13 Study Guide Gases Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 13 Study Guide Gases Answers
 - Highlighting and Note-Taking Chapter 13 Study Guide Gases Answers
 - Interactive Elements Chapter 13 Study Guide Gases Answers
- 8. Staying Engaged with Chapter 13 Study Guide Gases Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 13 Study Guide Gases Answers
- 9. Balancing eBooks and Physical Books Chapter 13 Study Guide Gases Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 13 Study Guide Gases Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 13 Study Guide Gases Answers
 - Setting Reading Goals Chapter 13 Study Guide Gases Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 13 Study Guide Gases Answers
 - Fact-Checking eBook Content of Chapter 13 Study Guide Gases Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
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