Chapter 10 Chemical Reactions

Chemical changes and reactions between compounds

Chemical formula revisited:

types of elements and ratio making up compound empirical formula: simplest whole number ratio of elements in a compound

> -general description of how to assemble the compound

ionic-exists as separate ions

-empirical formula is the right description

-simple mixture of ions Na⁺ and Cl⁻ ions

molecular (covalent)-exists as a particular

bonded entity

-empirical formula tells what it's made of

-does not describe how it is bonded

molecular formula: identifies actual numbers of atoms
in a molecule-discrete units of atoms
-not appropriate for ionic – array of separate ions
-molecules described by molecular formula
-you know a formula is molecular if the numbers
are not the simplest whole number ratio
glucose: C₆H₁₂O₆ molecular since divisible by 6

empirical - CH2O

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Arshad Iqbal

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