

Periodic Trends

Florida Sunshine State Standard Benchmark: SC.A.2.4.5- Knows that elements are arranged into groups and families based on similarities in electron structure and that their physical and chemical properties can be predicted.

Overview:

Elements within the modern periodic table are arranged, by increasing atomic number, in rows (periods) and columns (groups or families) in order to convey trends in their physical and chemical properties. In this way, generalizations concerning groups of elements can provide useful knowledge as to characteristic behaviors. A true understanding of the rationale behind the arrangement can provide insight into the behavior pattern of each element, such as its tendency to form a particular type of bond, react with a specific element, or change state.

Atomic Number		Atomic Number																Atomic Number	
1 2		3 4 5 6 7 8 9 10 11 12 13 14 15 16 17 18																19 20	
1s 2s		2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d 7p																3s 3p 4s	
1A 2A		3A 4A 5A 6A 7A 8A																1A 2A	
1	H																	2	He
2	Li	Be											B	C	N	O	F	Ne	
3	Na	Mg											Al	Si	P	S	Cl	Ar	
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	
5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	
6	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn	
7	Fr	Ra	Ac	Unk	Unp	Unh	Uns	Uno	Una	Uub									
Rare Earth Elements		Metals																Nonmetals	
Lanthanide Series		57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu			
Actinide Series		89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr			

The activity that follows provides students with the opportunity to compare and contrast properties of elements located throughout the periodic table as it details its organization.

Time Frame: 1 hour of class time, variable amount of research time

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Enjoying the Song of Appearance: An Emotional Symphony within **Ca Standards Periodic Table Trends Lesson Plan**

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