

Extension:

Osmotic pressure

Get the Gizmo ready:

- Select the STUDY SOLUTE tab.
- Set the Air temp. to 25 °C.



**Introduction:** Osmosis is the flow of water across a semipermeable membrane from a region of lower solute concentration to a region of higher solute concentration. The membrane allows water, but not solute, to go through. Like boiling and freezing, osmosis is a colligative property.

**Question:** How do different solutes affect osmosis?

1. Measure: Adding pressure to a solution can prevent water from entering the solution via osmosis. The osmotic pressure of the solution is equal to the pressure required to stop osmosis. A higher osmotic pressure indicates a more highly saturated solution.

Select **Sucrose** and click **Record**. What is the osmotic pressure of sucrose?

24.09

2. Predict: Based on the osmotic pressure of sucrose, predict what the osmotic pressures will be for the other solutes:

Sodium chloride: 28.02

Calcium chloride: 32.02

Potassium chloride: 28.02

3. Gather data: Use the Gizmo to measure the osmotic pressure of each solute. (Note: The Air temp. must be set to 25 °C to collect data on osmotic pressure.) Write the results below.

Sodium chloride: 48.18

Calcium chloride: 72.08

Potassium chloride: 48.18

4. Analyze: How close were your predictions? higher than I anticipated

5. Apply: Eating a diet rich in sodium can increase the risk of high blood pressure. Why would excessive amounts of sodium in the blood lead to higher blood pressure?

Since there would be an excess of salt in your system, more water would want to cross the membrane into the bloodstream which could cause greater pressure.

6. Summarize: What do all the colligative properties you have studied have in common?

They all have an impact on pressure, freezing point, and boiling point. The quantity of particles dissolved in the solution has an impact on them.

# Colligative Properties Virtual Lab Answers Key

**M Lipman**



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