

Compounds and Their Bonds

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A. Ionic Transfer of Electrons

Element	Atomic #	1. Electron Configuration of Atom	2. Electron-Dot Symbol	3. Loss or Gain	4. Electron Configuration of Ion	5. Ionic Charge	6. Symbol of Ion	7. Name of Ion
Sodium	11	$1s^2 2s^2 2p^6 3s^1$	$\cdot \text{Na} \cdot$	Loss 1 e ⁻	$1s^2 2s^2 2p^6$	1 ⁺	Na^+	Sodium ion
Nitrogen	7	$1s^2 2s^2 2p^3$	$\cdot \ddot{\text{N}} \cdot$	Gain 3 e ⁻	$1s^2 2s^2 2p^6$	3 ⁻	N^{3-}	Nitride ion
Aluminum	13	$1s^2 2s^2 2p^6 3s^2 3p^1$	$\cdot \ddot{\text{Al}} \cdot$	Loss 3 e ⁻	$1s^2 2s^2 2p^6$	3 ⁺	Al^{3+}	Aluminum ion
Chlorine	17	$1s^2 2s^2 2p^6 3s^2 3p^5$	$\cdot \ddot{\text{Cl}} \cdot$	Gain 1 e ⁻	$1s^2 2s^2 2p^6 3s^2 3p^6$	1 ⁻	Cl^-	Chloride ion
Calcium	20	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$	$\text{Ca} \cdot$	Loss 2 e ⁻	$1s^2 2s^2 2p^6 3s^2 3p^6$	2 ⁺	Ca^{2+}	Calcium ion
Oxygen	8	$1s^2 2s^2 2p^4$	$\cdot \ddot{\text{O}} \cdot$	Gain 2 e ⁻	$1s^2 2s^2 2p^6$	2 ⁻	O^{2-}	Oxide ion

B. Ionic Compounds and Formulas

1. Physical properties

Compound	Appearance	Density	Melting Point
Sodium chloride, NaCl	White, solid, crystalline	2.16 g/cm ³	801°C

2. Formulas of ionic compounds

Name	Positive Ion	Negative Ion	Formula
Lithium iodide	Li^{+1}	I^{-1}	LiI
Aluminum oxide	Al^{3+}	O^{2-}	Al_2O_3
Calcium sulfide	Ca^{2+}	S^{2-}	CaS
Magnesium bromide	Mg^{2+}	Br^{-1}	MgBr_2
Potassium nitride	K^{+1}	N^{3-}	K_3N
Sodium fluoride	Na^{+1}	F^{-1}	NaF

3. Names of ionic compounds

Formula	Positive Ion	Negative Ion	Name
K_2S	K^{+1}	S^{2-}	Potassium sulfide
BaF_2	Ba^{2+}	F^{-1}	Barium fluoride
MgO	Mg^{2+}	O^{2-}	Magnesium oxide
Na_3N	Na^{+1}	N^{3-}	Sodium nitride
AlCl_3	Al^{3+}	Cl^{-1}	Aluminum chloride
Mg_3P_2	Mg^{2+}	P^{3-}	Magnesium phosphide

Compounds Their Formulas Lab 7 Answers

KJ Lindholm-Leary



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