



Diffusion and Osmosis

OBJECTIVES:

- To gain a better understanding of diffusion and osmosis.
- To understand these terms: diffusion, osmosis, concentration gradient, Brownian motion, hypotonic, hypertonic, isotonic, hypothesis, selectively permeable, prediction, control, independent variable, dependent variable.
- To further your understanding of experimental design and practice hypothesis testing.

INTRODUCTION

This lab will introduce you to diffusion and osmosis, especially as they relate to cell membranes.

In 1827, an English botanist named Robert Brown noticed that pollen grains suspended in water move around randomly under a microscope. The smaller the particle is, the faster it jiggles. Atoms and small molecules bounce around in liquids and gases. The hotter the temperature, the faster they jiggle. (Temperature is actually a measure of kinetic energy of jiggling atoms and molecules.) The constant random motion of particles in gases and liquids is called **Brownian motion**. Brownian motion can be seen by dropping India ink into water on a glass slide and focusing on it under a microscope. (For example, search for "Brownian motion" in Youtube.com.)

Brownian motion causes molecules to bounce around even in water that looks still to the naked eye. The constant bouncing mixes everything up at an atomic level. This explains why food dye, when added to water, spreads out until the dye molecules are evenly distributed in the liquid (see Figure 1). **Diffusion** is the spontaneous movement of solutes from an area of high concentration to an area of low concentration. **Solutes** are atoms and molecules that are suspended in a liquid (or a gas). The liquid that solutes are suspended in is called the **solvent**. In biological systems, the solvent is usually water. Solutes can be any atoms and molecules in the water, such as salts, sugars, amino acids and even proteins. Once solutes are evenly spread out, they are in equilibrium, meaning that there is no net movement of solutes in any direction even though they are still bouncing around randomly due to Brownian motion.

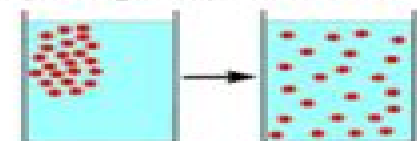


Figure 1

Diffusion is a passive process. No energy input is required. Diffusion only requires a **concentration gradient**, which is a difference in concentration over a distance. A sugar cube dropped into coffee without stirring will dissolve and spread throughout the entire cup of coffee, but some time is required to reach equilibrium. (Diffusion is slow compared to bulk stirring, such as stirring the coffee with a spoon.) Though slow, diffusion in liquids is fast enough over short distances to drive the passive transport of nutrients and wastes across biological membranes. Biological membranes are **selectively permeable (semipermeable)**, meaning they let some solutes through but do not allow other solutes to go through. Transport proteins are required to move ions and polar molecules into and out of cells. Transport proteins are very specific about which solutes they will transport across membranes. If no transport protein is present for a certain charged or polar solute, that solute cannot be moved into and out of cells by diffusion. On the other hand, small nonpolar gases such as CO_2 and O_2 readily diffuse through the phospholipid bilayer of biological membranes.

Water can also diffuse down its concentration gradient. Because water is the solvent in biological systems, it can be a bit confusing to consider, and so the diffusion of water is given a special name.

Osmosis is the diffusion of water from an area of high water concentration to an area of low water concentration. Higher water concentration means lower solute concentration (purer water). Lower water concentration means higher solutes (less pure water). Water often passes through biological membranes faster than solutes and osmosis is the driving force for moving water into and out of cells across biological membranes.

Plant root cells contain ions, sugars, amino acids, proteins and other solutes. The solute concentration is higher inside root cells compared to the dilute watery environment outside cells. From the solvent point

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