

You need to know the content of this sheet. 100%

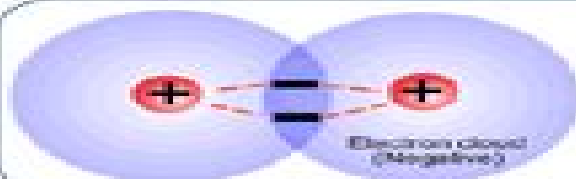
100% Sheet COVALENT BONDING



What's the
BIG
idea?

Matter

All material in the Universe is made of very small particles.



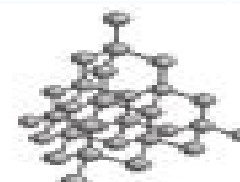
A Covalent bond forms between **2 non-metals** where each atom **shares** its electrons in order to obtain a full outer shell

Simple (small) covalent molecules are mostly gases at room temperature (except water) because they have **weak (easy to break) bonds** between the molecules. The **STRONG COVALENT** bonds **DO NOT BREAK**



Graphite

1. 3 strong covalent bonds
2. 1 weak intermolecular bond between the layers
3. layers can slide over each other
4. Free (Delocalised) electrons between the layers so it can conduct electricity



Diamond

1. 4 strong covalent bonds so high melting point
2. No weak bonds so it is very strong and hard
3. No free electrons or charges so it cannot conduct electricity

Other structures of carbon can be man-made. Nanotubes are made by folding a sheet of graphene (which is a single sheet of graphite) these also have delocalised electrons



fullerene



nanotube



graphene

You **MUST** be able to draw dot and cross diagrams for the molecules of hydrogen, chlorine, oxygen, nitrogen, hydrogen chloride, water, ammonia and methane



Covalent Bonding Core Teaching Resources Test Answer

Yicheng Fang



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