

Name _____ Date _____ Class _____

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Covalent Bonding

Section 9.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	p bond
---------------	----------	------------	------------	--------

When sharing of electrons occurs, the attachment between atoms that results is called a(n) (1) covalent bond. When such an attachment is formed, bond dissociation energy is released, and the process is (2) exothermic. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) (3) molecule. If the electrons shared are contained between the two atoms, the attachment is called a(n) (4) sigma bond. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) (5) p bond.

In your textbook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- In what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
 - a single atom
 - (b) an molecules containing two atoms
 - a molecules containing three atoms
 - a molecules containing four atoms
- How many electrons are shared in a double covalent bond?
 - none
 - one
 - two
 - (d) four
- Bond lengths in the distance between
 - two molecules of the same substance
 - (c) the nuclei of two attached atoms
 - the electrons in two attached atoms
 - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
 - (a) the shorter the bond, the stronger the bond
 - the shorter the bond, the weaker the bond
 - the shorter the bond, the fewer the electrons in it
 - the shorter the bond, the lower the bond dissociation energy

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Chemistry: Matter and Change • Chapter 9

49

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CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Section 9.2 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write *true* or *false*.

- false Binary molecular compounds are generally composed of a metal and a nonmetal.
- true The second element in the formula of a binary compound is named using the suffix *-ite*.
- false The prefix *tri-* indicates three atoms.
- true The prefix *kilo-* indicates six atoms.
- true In naming the first element in a formula, the prefix *mono-* is not used.
- true For binary acids, the hydrogen part of the compound is named using the prefix *hydro-*.
- false An crystal contains only two elements.
- false If the name of the cation of an ionic salt is only the salt name contains the suffix *-ate*.

In your textbook, read about naming molecular compounds and oxoacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A	Column B
<u>c</u> H_2CO	a. hydrobromic acid
<u>i</u> NO	b. nitrogen trioxide
<u>g</u> H_2CO_3	c. carbon monoxide
<u>e</u> NO_2	d. nitrous acid
<u>b</u> N_2O_5	e. oxoacids
<u>d</u> HNO_3	f. nitric acid
<u>f</u> HNO_2	g. carbonic acid
<u>a</u> HBr	h. bromic acid
<u>h</u> HBrO_3	i. carbon dioxide

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Chemistry: Matter and Change • Chapter 9

Chemistry Study Guide Content Mastery Teacher

**Jhoni Warmansyah, Artha Nesa
Chandra, M. Haviz, Isra Nurmai
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1. Understanding the eBook Chemistry Study Guide Content Mastery Teacher
 - The Rise of Digital Reading Chemistry Study Guide Content Mastery Teacher
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Content Mastery Teacher
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Content Mastery Teacher
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Content Mastery Teacher
 - Personalized Recommendations
 - Chemistry Study Guide Content Mastery Teacher User Reviews and Ratings
 - Chemistry Study Guide Content Mastery Teacher and Bestseller Lists
5. Accessing Chemistry Study Guide Content Mastery Teacher Free and Paid eBooks
 - Chemistry Study Guide Content Mastery Teacher Public Domain eBooks
 - Chemistry Study Guide Content Mastery Teacher eBook Subscription Services
 - Chemistry Study Guide Content Mastery Teacher Budget-Friendly Options
6. Navigating Chemistry Study Guide Content Mastery Teacher eBook Formats

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- Chemistry Study Guide Content Mastery Teacher Compatibility with Devices
- Chemistry Study Guide Content Mastery Teacher Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Chemistry Study Guide Content Mastery Teacher
 - Highlighting and Note-Taking Chemistry Study Guide Content Mastery Teacher
 - Interactive Elements Chemistry Study Guide Content Mastery Teacher
- 8. Staying Engaged with Chemistry Study Guide Content Mastery Teacher
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Content Mastery Teacher
- 9. Balancing eBooks and Physical Books Chemistry Study Guide Content Mastery Teacher
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Content Mastery Teacher
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Study Guide Content Mastery Teacher
 - Setting Reading Goals Chemistry Study Guide Content Mastery Teacher
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Study Guide Content Mastery Teacher
 - Fact-Checking eBook Content of Chemistry Study Guide Content Mastery Teacher
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
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