

## Stuff I Should Know (Page 2)

### Complex Ions & Common Ligands

| Ligands      | polar molecules & anions                                                                                   | $\text{NH}_3$ , $\text{H}_2\text{O}$ , $\text{OH}^-$ , $\text{CN}^-$ , $\text{Cl}^-$                                                                              | Odd example:<br>$\text{Fe}^{3+} + \text{SCN}^- \rightleftharpoons \text{FeSCN}^{2+}$                           |
|--------------|------------------------------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------------|
| Central ions | transition metals and $\text{Al}^{3+}$                                                                     | $\text{Ag}^+$ , $\text{Cu}^{2+}$ , $\text{Ni}^{2+}$ , $\text{Zn}^{2+}$ , etc. & $\text{Al}^{3+}$                                                                  |                                                                                                                |
| Examples     | usually twice the number of ligands as the charge on the central ion.<br>Key Words: "excess, concentrated" | $\text{Ag}(\text{CN})_2^-$ , $\text{Cu}(\text{NH}_3)_4^{2+}$ ,<br>$\text{Ni}(\text{OH})_4^{2-}$ , $\text{Zn}(\text{NH}_3)_4^{2+}$ ,<br>$\text{Al}(\text{OH})_4^-$ | Reaction with Acid:<br>$\text{Cu}(\text{NH}_3)_4^{2+} + \text{H}^+ \rightarrow \text{Cu}^{2+} + \text{NH}_4^+$ |

### Organic Chemistry & Functional Groups

| alkanes<br>$\text{C}_n\text{H}_{2n+2}$                                                         | alkenes<br>$\text{C}_n\text{H}_{2n}$                                                         | alkynes<br>$\text{C}_n\text{H}_{2n-2}$                                               | aromatics (benzene)<br>$\text{C}_6\text{H}_6$                                          |
|------------------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------|
| alcohol<br><br>$\text{R}-\text{OH}$                                                            | aldehyde<br><br>$\text{R}-\overset{\text{O}}{\underset{\text{  }}{\text{C}}}-\text{H}$       | ketone<br><br>$\text{R}-\overset{\text{O}}{\underset{\text{  }}{\text{C}}}-\text{R}$ | ether<br><br>$\text{R}-\text{O}-\text{R}$                                              |
| carboxylic acid<br><br>$\text{R}-\overset{\text{O}}{\underset{\text{  }}{\text{C}}}-\text{OH}$ | ester<br><br>$\text{R}-\overset{\text{O}}{\underset{\text{  }}{\text{C}}}-\text{O}-\text{R}$ | amine<br><br>$\text{R}-\text{NH}_2$                                                  | amide<br><br>$\text{R}-\overset{\text{O}}{\underset{\text{  }}{\text{C}}}-\text{NH}_2$ |
| Substituted benzene:                                                                           | ortho = 1,2                                                                                  | meta = 1,3                                                                           | para = 1,4                                                                             |

### nuclear chem

|                                      |
|--------------------------------------|
| alpha<br>${}^4_2\text{He}$           |
| beta electron<br>${}^0_{-1}\text{e}$ |
| neutron<br>${}^1_0\text{n}$          |
| positron<br>${}^0_{+1}\text{e}$      |

### $\Delta H$ $\Delta S$ Spont?

|                                                        |                   |              |
|--------------------------------------------------------|-------------------|--------------|
| -                                                      | +                 | at all temps |
| +                                                      | +                 | high temps   |
| -                                                      | -                 | low temps    |
| +                                                      | -                 | no temps     |
| Note: $\Delta S$ in J<br>$\Delta G$ & $\Delta H$ in kJ |                   |              |
| <b><math>K_{sp}</math> &amp; Solubility, s</b>         |                   |              |
| 1:1                                                    | $K_{sp} = s^2$    |              |
| 1:2                                                    | $K_{sp} = 4s^3$   |              |
| 1:3                                                    | $K_{sp} = 27s^4$  |              |
| 2:3                                                    | $K_{sp} = 108s^5$ |              |

### Lewis Acids & Bases

$\text{BF}_3 + \text{NH}_3 \rightarrow \text{BF}_3\text{NH}_3$   
acid anhydrides (oxides of nonmetals,  $\text{CO}_2$ )  
basic anhydrides (oxides of metals,  $\text{MgO}$ )  
 $\text{MgO} + \text{CO}_2 \rightarrow \text{MgCO}_3$   
decomposition reactions:  $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$   
Strange Examples:  $\text{P}_2\text{O}_5 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_4$

### Strange Ions: (nitride, $\text{N}^{3-}$ ) (hydride, $\text{H}^-$ )



### Flame Test Colors

|                          |
|--------------------------|
| Barium – green           |
| Sodium – yellow          |
| Copper – blue (w/ green) |
| Potassium – lavender     |
| Strontium – red          |
| Lithium – red            |
| Calcium – orange         |

### Quantum Numbers

| n     | 1, 2, 3, ...                  |
|-------|-------------------------------|
| l     | 0 ... (n-1)                   |
| m     | -l ... +l                     |
| $m_l$ | +l, ..., -l                   |
| l     | 0 = s, 1 = p,<br>2 = d, 3 = f |

### Writing Lewis Structures

hint: use one valence electron to connect F's or Cl's then determine lone pairs (Ex:  $\text{XeF}_4$ )

### Product-Favored (Spontaneous) Reactions

$$\Delta G < 0 \quad E^\circ > 0 \quad K_{eq} > 1$$

### Properties Indicate Strength of Intermolecular Forces (IMF's)

| IMF | BP | FP | $H_{vap}$ | $H_{fus}$ | VP |
|-----|----|----|-----------|-----------|----|
| IMF | BP | FP | $H_{vap}$ | $H_{fus}$ | VP |

### Orders of Reactions & Graphs That Give Straight Lines

| 0 <sup>th</sup> Order | 1 <sup>st</sup> Order    | 2 <sup>nd</sup> Order   |
|-----------------------|--------------------------|-------------------------|
| $[\text{R}]$ vs. Time | $\ln[\text{R}]$ vs. Time | $1/[\text{R}]$ vs. Time |
| slope = -k            | slope = -k               | slope = -k              |

### Electrochemical Cells

| anode           | cathode          |
|-----------------|------------------|
| oxidation       | reduction        |
| - side          | + side           |
| lower $E^\circ$ | higher $E^\circ$ |
| $e^-$ leave     | $e^-$ enter      |

### Bond Orders

| bond   | B.O. | $\sigma$        |
|--------|------|-----------------|
| single | 1    | $\sigma$        |
| double | 2    | $\sigma + \pi$  |
| triple | 3    | $\sigma + 2\pi$ |

### $\text{SN}$ & hybridization & shape

| Steric Number | hybridization                  | basic shape          |
|---------------|--------------------------------|----------------------|
| 1             | s                              | —                    |
| 2             | sp                             | linear               |
| 3             | sp <sup>2</sup>                | $\Delta$ planar      |
| 4             | sp <sup>3</sup>                | tetrahedral          |
| 5             | sp <sup>3</sup> d              | $\Delta$ bipyramidal |
| 6             | sp <sup>3</sup> d <sup>2</sup> | octahedral           |

### IMF's

|                  |                                                                                           |
|------------------|-------------------------------------------------------------------------------------------|
| London           | nonpolar molecules, ex: $\text{CH}_4$ , He                                                |
| dipole-dipole    | polar molecules, ex: $\text{H}_2\text{S}$ , $\text{SO}_2$                                 |
| hydrogen bonding | H-F, H-O-, H-N-, $\text{NH}_3$ , $\text{H}_2\text{O}$<br>amines and alcohols              |
| metallic         | metals, Ag, Pb                                                                            |
| ionic            | salts, $\text{NaCl}$ , $\text{CaCO}_3$<br>(Note: "ates" contain covalent bonds)           |
| covalent network | C (graphite), C (diamond), $\text{SiO}_2$ , WC,<br>Si, SiC (Note: graphite = London, too) |

### Activity of Metals (Four Groups)

| Metals        | React with...                                                                                                                                                 |
|---------------|---------------------------------------------------------------------------------------------------------------------------------------------------------------|
| Groups I & II | $\text{H}_2\text{O}$ ex: $\text{Li} + \text{H}_2\text{O} \rightarrow \text{Li}^+ + \text{OH}^- + \text{H}_2$                                                  |
| all others    | Non-oxidizing Acid, ex: $\text{HCl}$<br>$\text{Zn} + 2\text{HCl} \rightarrow \text{H}_2 + \text{ZnCl}_2$                                                      |
| Cu, Ag, Hg    | Oxidizing Acid, $\text{HNO}_3$ or $\text{H}_2\text{SO}_4$ (conc.)<br>$\text{Cu} + \text{HNO}_3 \rightarrow \text{NO}_2 + \text{H}_2\text{O} + \text{Cu}^{2+}$ |
| Au, Pt, Ir    | Aqua Regia ( $\text{HNO}_3$ + $\text{HCl}$ )                                                                                                                  |

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**JG Myers**



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