

Opgave 1 Uranium

- a. $^{235}_{92}\text{U} + ^1_0\text{n} \rightarrow 3^1_0\text{n} + ^{142}_{56}\text{Ba} + ^{91}_{36}\text{Kr}$
 dus $p = 36$, $n = 55$; symbool Kr
- b. $251 \text{ g U} \hat{=} 251 / 235 = 1,068 \text{ mol U} \hat{=} 1,068 \text{ mol Ba} \hat{=} 1,068 \times 142 = 152 \text{ g Ba}$

Opgave 2 Perovskiet

- a. Het zout bestaat uit CH_3NH_3^+ , 3 keer I^- en lood met een onbekende lading. De totale negatieve lading is $3-$.
 De totale lading in een zout is altijd 0, daarom hebben we hier Pb^{2+} . Pb^{2+} en CH_3NH_3^+ maakt samen een lading van $3+$.
- b. De totale massa van $\text{CH}_3\text{NH}_3\text{PbI}_3 = 12,01 + 6 \times 1,008 + 14,01 + 207,2 + 3 \times 126,9 = 620,0 \text{ g/mol}$
 $207,2 / 620,0 \times 100 \% = 33,4 \text{ mol\% lood}$
- c. Tin staat boven lood/in dezelfde groep als lood in het periodiek systeem. Omdat tin en lood in dezelfde groep staan hebben ze vergelijkbare eigenschappen. Daarom zou tin zoals lood kunnen werken in een perovskietachtige stof.
- d. $4,2 \times 10^3 \text{ gram} / 207,2 = 20,27 \text{ mol Pb}$
 In 1 mol perovskiet komt 1 mol Pb voor, dus kun je ook 20,27 mol perovskiet maken.
 $20,27 \text{ mol} \times 620,0 \text{ g/mol} = 1,26 \cdot 10^4 \text{ gram perovskiet}$
- e. $1,26 \cdot 10^4 / 4,1 = 3,07 \cdot 10^3 \text{ cm}^3$. Dat komt overeen met $3,1 \cdot 10^{-3} \text{ m}^3$ perovskiet.
 Als je hebt gerekend met het in de opgave gegeven foute antwoord van $1,0 \cdot 10^4 \text{ gram}$, kom je op $2,4 \cdot 10^{-3} \text{ m}^3$ perovskiet.
- f. $30 \times 3000 = 90000 \text{ kWh}$ is nodig; $90000 / 220 = 409$ zonnepanelen zijn nodig.
 $409 \times 1,8 \text{ m}^2 = 736 \text{ m}^2$ zonnepanelen.
 $3,1 \cdot 10^{-3} / 736 = 4,2 \cdot 10^{-6} \text{ m} = 4,2 \cdot 10^{-3} \text{ mm}$. De perovskietlaag van een zonnepaneel zou dan $4,2 \cdot 10^{-3} \text{ mm}$ dun zijn.
 Als je hebt gerekend met het in de opgave gegeven foute antwoord van $2,0 \cdot 10^{-3} \text{ m}^3$ perovskiet kom je op $2,7 \cdot 10^{-3} \text{ mm}$.

Opgave 3 TGG-waarde

- a. $V(\text{lucht}) = 200 \text{ m}^3$ en TGG = $15 \text{ mg/m}^3 \Rightarrow$ Er mag dus $200 \times 15 = 3,0 \cdot 10^3 \text{ mg H}_2\text{S}$ vrijkomen.
 $V(\text{lucht}) = 200 \text{ m}^3$ en $\rho = 1,3 \text{ kg/m}^3 \Rightarrow m(\text{lucht}) = 1,3 \times 200 = 260 \text{ kg}$.
 Reukgrens is $6,8 \text{ mg/kg lucht}$, dus (bij de reukgrens) in 260 kg zit dan $6,8 \times 260 = 1,8 \cdot 10^3 \text{ mg H}_2\text{S}$.
 Bij de reukgrens is de TGG dus nog niet overschreden.
- b. Reukgrens is $6,8 \text{ mg/kg lucht}$.
 - berekening $V(\text{H}_2\text{S}) = 6,8 \cdot 10^{-6} / 1,43 = 4,8 \cdot 10^{-6} \text{ m}^3$
 - berekening $V(\text{lucht}) = 1,0 / 1,3 = 0,77 \text{ m}^3$

$$\text{volume} - \text{ppm} = \frac{V_{\text{gas}}}{V_{\text{gas+lucht}}} \cdot 10^6 = \frac{4,8 \cdot 10^{-6}}{0,77} \cdot 10^6 = 6,2 \text{ vol} - \text{ppm}$$
- c. $2 \text{ H}_2\text{S} + 3 \text{ O}_2 \rightarrow 2 \text{ SO}_2 + 2 \text{ H}_2\text{O}$
- d. In $1,2 \text{ ton}$ zit $1,2 \cdot 10^6 \times 120 \cdot 10^{-6} = 144 \text{ g H}_2\text{S}$
 $n(\text{H}_2\text{S}) = 144 / 34,076 = 4,226 \text{ mol H}_2\text{S}$
 $n(\text{SO}_2) = 2/2 \times 4,226 = 4,226 \text{ mol SO}_2 \hat{=} 4,226 \times 24,5 = 1,0 \cdot 10^2 \text{ dm}^3 \text{ SO}_2$

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