

Concentration and Molarity PhET Labs

Name: _____

Procedure: https://phet.colorado.edu/sims/html/concentration/latest/concentration_en.html

Part 1: Dissolution and Saturation

Take some time to play and familiarize yourself with the simulation. Click on everything. Move all the sliders. Notice what happens to the concentration as solid solute is added and when evaporation occurs.



Concentration

- How does the concentration change as solid solute is added? concentration increases
 - How does the concentration change as additional water is added? concentration decreases
 - How does the concentration change as evaporation occurs? concentration increases
 - How do you know when a solution is saturated? concentration does not change (it also shows on screen)
 - When a solution is saturated, and additional solid solute is added, what happens? solid does not dissolve and settles at bottom
 - Why do you think this is? solvent can't hold any more solute
 - How does adding this additional solute change the concentration of this saturated solution? decreases concentration
 - How does evaporation change the concentration of a saturated solution? no change- solution is already saturated
- Part 2: Concentrated Solutions (same website as part 1 just click on the Solution option)
- Adding a concentrated solution ... describe a way to determine the concentration of the solution in the spigot. Write your plan here: empty the container and add only one solution from the dropper and then use the sensor to measure concentration



- Using your plan ... how might you get that concentrated solution to become saturated? you can evaporate
- Does your plan work for all the other solutions too? yes Why? / Why Not? evaporation reduces the amount of water (or solute) so concentration increases



Part 3: Molarity https://phet.colorado.edu/sims/html/molarity/latest/molarity_en.html

Molarity (M) = $\frac{\text{moles of solute (mol)}}{\text{volume of solution (L)}}$

Molarity is moles per liter, that is, how many moles of solute (entire salt) is dissolved per liter of solution.

First, determine the saturation concentration of each of the solutions, that is, how concentrated can you get each solution before the solution is saturated. If you can't determine the concentration using the simulation "Molarity", try using the simulation "Concentration" (You will use this information again in Part 5, if your instructor requires it)



Molarity

Cobalt (II) nitrate	<u>5.640 M</u>	Potassium chromate	<u>3.350 M</u>
Cobalt chloride	<u>4.330 M</u>	Nickel (II) chloride	<u>5.210 M</u>
Potassium dichromate	<u>0.510 M</u>	Copper sulfate	<u>1.380 M</u>
Gold (III) chloride	<u>2.30 M</u>	Potassium permanganate	<u>0.480 M</u>

Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises

C Cleary



Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises:

This book delves into Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises. Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
- Chapter 2: Essential Elements of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
- Chapter 3: Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises in Everyday Life
- Chapter 4: Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises. This chapter will explore what Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises is, why Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises is vital, and how to effectively learn about Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises.
3. In chapter 2, this book will delve into the foundational concepts of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises. The second chapter will elucidate the essential principles that must be understood to grasp Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises in its entirety.
4. In chapter 3, the author will examine the practical applications of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises in daily life. The third chapter will showcase real-world examples of how Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises in specific contexts. The fourth chapter will explore how Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises. This chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Conclusion Questions And Calculations

Concentration And Molarity Post Lab Exercises.

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