

CHEMISTRY



Periodic Table of Elements

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K 39.098	Ca 40.078	Sc 44.956	Ti 47.88	V 50.942	Cr 52.00	Mn 54.938	Fe 55.845	Co 58.933	Ni 58.69	Cu 63.546	Zn 65.38	Ga 69.723	Ge 72.64	As 74.922	Se 78.96	Br 79.904	Kr 83.80																																																																
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Rb 85.468	Sr 87.62	Y 88.906	Zr 91.224	Nb 92.906	Mo 95.94	Tc 98.906	Ru 101.07	Rh 101.07	Pd 106.36	Ag 107.868	Cd 112.411	In 114.818	Sn 118.710	Sb 121.757	Te 127.60	I 126.905	Xe 131.29	Cs 132.905	Ba 137.327	La 138.905	Ce 140.12	Pr 140.908	Nd 144.24	Pm 144.913	Sm 150.36	Eu 151.964	Gd 157.25	Tb 158.925	Dy 162.50	Ho 164.930	Er 167.259	Tm 168.930	Yb 173.054	Lu 174.967	Hf 178.49	Ta 180.948	W 183.84	Re 186.207	Os 190.23	Ir 192.222	Pt 195.084	Au 196.967	Hg 200.59	Tl 204.38	Pb 207.2	Bi 208.980	Po 209	At 210	Rn 222	Fr 223	Ra 226	Ac 227	Th 232.038	Pa 231.036	U 238.029	Np 237.048	Pu 244.064	Am 243.061	Cm 247.070	Bk 247.070	Cf 251.08	Es 252.083	Fm 257.10	Md 258.10	No 259.10	Lr 262.10															
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Rf 261.10	Db 262.10	Sg 266.10	Bh 264.10	Hs 277.10	Mt 268.10	Ds 271.10	Rg 272.10	Cn 285.10	Nh 286.10	Fl 289.10	Mc 288.10	Lv 293.10	Ts 294.10	Og 294.10																																																																			

1	ATOMIC NUMBER
H	SYMBOL
1.008	ATOMIC WEIGHT

- Elemental forms:**
 - Monoatomic:** Noble gases: helium, neon, argon, krypton, xenon, and radon; nonreactive
 - Diatomic:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; **Liquid:** bromine; **Solid:** iodine
 - Molecular solids:** Sulfur
 - Covalent solids:** Carbon (diamond, graphite)
 - Metals:** Most common type of element; high coordination numbers; very dense
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options
 - Ionic:** Electron exchange
 - Covalent:** Shared electrons
 - Polar covalent: Unequal sharing
 - Ion complex:** Central ion + Solvent
 - Transition metal compounds:**
 - Coordination compounds: Non-bonding ligands
 - Organometallics: Metal-C bond
 - Complex minerals:** Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
 - Family of elements:** Column in periodic table
 - The addition of d-orbitals produces an expanded octet

Periodic Trends

- Chemistry determined by the behavior of valence electrons
- Atoms or ions get larger as you go down the column; electron shells are added
- Cation and anions (same charge) decrease in size as the atomic number (Z) increases

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to ionize; electron screening lessens Z -effect on orbital energy
- Atomic and ionic radii:** Not real properties; derive from structural data; divide bond into radii using a consistent scheme
 - Ionic radius:** Depends on charge, counter ion, and coordination (crystal type)
 - Atomic radius:** Based on diatomic gas data or metallic solid structural data

Development of the Periodic Table

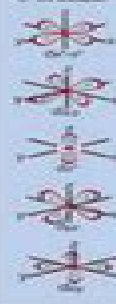
- Missing elements were predicted using periodicity
- Column 1:** Alkali metals
- Column 2:** Alkaline earth metals
- First transition series:** Titanium-Copper
- Column 17:** Halogens: fluorine, chlorine, bromine, iodine, and astatine
- Column 18:** Rare gases: helium, neon, argon, krypton, xenon, and radon
- Actinides:** $Z = 90$ or above; radioactive
- Lanthanides:** f-subshell elements; starts with $Z = 58$



Practical Application

- The colors of fireworks depend on the elements in the explosive

d-subshell 5 orbitals



p-subshell 3 orbitals



s-subshell 1 orbital



Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an anion of charge $-q$, $Z = +q$ electrons; for a cation of charge $+q$, $Z = -q$ electrons
- Mass number (A):** $A = Z + N$, where N = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same Z , different A ; an element can have one or more isotopes; any sample contains a number of isotopes; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element)



Atomic Quantum Numbers & Orbitals

- n : principle**
 - n : magnetic (orbital direction)
 - n : electron spin
- For each n , the possible l values are $0, 1, \dots, n-1$
- For each l , the possible m_l values are $-l, \dots, 0, \dots, +l$
- m_s has two possible values: $+\frac{1}{2}$ and $-\frac{1}{2}$ (spin up, spin down)
- Each value of n denotes a "shell" in the atomic structure; each l denotes a subshell
 - s-orbital: $l = 0$, 1 type
 - d-orbital: $l = 2$, 5 types
 - p-orbital: $l = 1$, 3 types
 - f-orbital: $l = 3$, 7 types

Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbau principle:

Aufbau Order for Filling Subshells



n	l	m_l	m_s	element
1	0	0	$+\frac{1}{2}$	H
1	0	0	$-\frac{1}{2}$	He
2	0	0	$+\frac{1}{2}$	Li
2	0	0	$-\frac{1}{2}$	Be
2	1	-1	$+\frac{1}{2}$	B
2	1	0	$+\frac{1}{2}$	C
2	1	1	$+\frac{1}{2}$	N
2	1	-1	$-\frac{1}{2}$	O
2	1	0	$-\frac{1}{2}$	F
2	1	1	$-\frac{1}{2}$	Ne

- Electron configuration:** Atomic orbital occupancy for an atom
- Pauli exclusion principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons: one with spin up, $m_s = +\frac{1}{2}$, and one with spin down, $m_s = -\frac{1}{2}$
- Hund's rule:** Electrons fill p, d, and f subshells in a manner to maximize the number of unpaired electrons (half-filled orbitals); spin is maximized
- Ionization potential (IP):** Energy required to remove electron from an atom or ion
 - First IP:** Removal of first valence electron
- Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale, 0 for rare gas

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