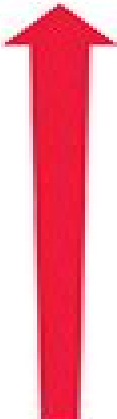


Acids and Bases

Acid				Base			
Strength of Acid	K _a	Name	Formula	Formula	Name	Strength of Base	
Strongest						Weakest	
	Large	Perchloric acid	HClO ₄	ClO ₄ ⁻	Perchlorate ion		
	3.2×10 ⁸	Hydroiodic acid		I ⁻	Iodide		
	1.0×10 ⁹	Hydrobromic acid	HBr	Br ⁻	Bromide		
	1.3×10 ⁹	Hydrochloric acid	HCl	Cl ⁻	Chloride		
	1.0×10 ³	Sulfuric acid	H ₂ SO ₄	HSO ₄ ⁻	Hydrogen sulfate ion		
	2.4×10 ¹	Nitric acid	HNO ₃	NO ₃ ⁻	Nitrate ion		
	-----	Hydronium ion	H ₃ O ⁺	H ₂ O	Water		
	5.4×10 ⁻²	Oxalic acid	HO ₂ C ₂ O ₂ H	HO ₂ C ₂ O ₂ ⁻	Hydrogen oxalate ion		
	1.3×10 ⁻²	Sulfurous acid	H ₂ SO ₃	HSO ₃ ⁻	Hydrogen sulfite ion		
	1.0×10 ⁻²	Hydrogen sulfate ion	HSO ₄ ⁻	SO ₄ ²⁻	Sulfate ion		
	7.1×10 ⁻³	Phosphoric acid	H ₃ PO ₄	H ₂ PO ₄ ⁻	Dihydrogen phosphate ion		
	7.2×10 ⁻⁴	Nitrous acid	HNO ₂	NO ₂ ⁻	Nitrite ion		
	6.6×10 ⁻⁴	Hydrofluoric acid	HF	F ⁻	Fluoride ion		
	1.8×10 ⁻⁴	Methanoic acid	HCO ₂ H	HCO ₂ ⁻	Methanoate ion		
	6.3×10 ⁻⁵	Benzoic acid	C ₆ H ₅ COOH	C ₆ H ₅ COO ⁻	Benzoate ion		
	5.4×10 ⁻⁵	Hydrogen oxalate ion	HO ₂ C ₂ O ₂ ⁻	O ₂ C ₂ O ₂ ²⁻	Oxalate ion		
	1.8×10 ⁻⁵	Ethanoic acid	CH ₃ COOH	CH ₃ COO ⁻	Ethanoate (acetate) ion		
	4.4×10 ⁻⁷	Carbonic acid	H ₂ CO ₃	HCO ₃ ⁻	Hydrogen carbonate ion		
	1.1×10 ⁻⁷	Hydrosulfuric acid	H ₂ S	HS ⁻	Hydrogen sulfide ion		
	6.3×10 ⁻⁸	Dihydrogen phosphate ion	H ₂ PO ₄ ⁻	HPO ₄ ²⁻	Hydrogen phosphate ion		
	6.2×10 ⁻⁸	Hydrogen sulfite ion	HS ⁻	SO ₃ ²⁻	Sulfite ion		
	2.9×10 ⁻⁸	Hypochlorous acid	HClO	ClO ⁻	Hypochlorite ion		
	6.2×10 ⁻¹⁰	Hydrocyanic acid	HCN	CN ⁻	Cyanide ion		
	5.8×10 ⁻¹⁰	Ammonium ion	NH ₄ ⁺	NH ₃	Ammonia		
	5.8×10 ⁻¹⁰	Boric acid	H ₃ BO ₃	H ₂ BO ₃ ²⁻	Dihydrogen borate		
	4.7×10 ⁻¹¹	Hydrogen carbonate ion	HCO ₃ ⁻	CO ₃ ²⁻	Carbonate ion		
	4.2×10 ⁻¹³	Hydrogen phosphate ion	HPO ₄ ²⁻	PO ₄ ³⁻	Phosphate ion		
	1.8×10 ⁻¹³	Dihydrogen borate ion	H ₂ BO ₃ ⁻	HBO ₃ ²⁻	Hydrogen borate ion		
	1.3×10 ⁻¹³	Hydrogen sulfide ion	HS ⁻	S ²⁻	Sulfide ion		
	1.6×10 ⁻¹⁴	Hydrogen borate ion	HBO ₃ ²⁻	BO ₃ ³⁻	Borate ion		
Weakest	-----	water	H ₂ O	OH ⁻	Hydroxide	Strongest	

Chemistry Acids And Bases Names

Neil D. Jespersen, Alison Hyslop



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Organic Chemistry T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-01-19 The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show

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Copeland, 1903 *The Encyclopedia Americana*, 1918

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