

# UNIT 1: MATTER & MEASUREMENT

## Experimental Design

**Independent Variable** = Manipulated condition  
**Dependent Variable** = Responding variable  
**Controlled Variable** = Must be same for all groups

As part of the trial, substances are mixed in 4 beakers to see how much acid reacts with the sodium hydroxide solution in each beaker.



**Independent Variable** = HCl Concentration (molarity)  
**Dependent Variable** = Reaction time  
**Controlled Variable** = Temperature, Hg sample size

## Physical vs. Chemical

**Physical:**

- \* Substance retains its properties
- \* Chemical formulas do not change
- \* Usually Phase Change

Examples:



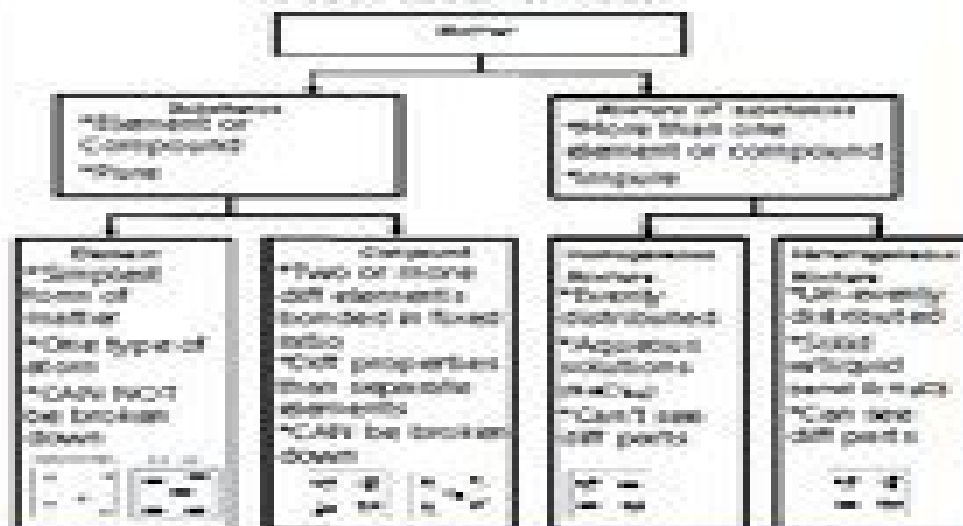
**Chemical:**

- \* Properties do not change
- \* Chemical formulas change
- \* rust, burn, oxidize, combust, corrode

Examples:



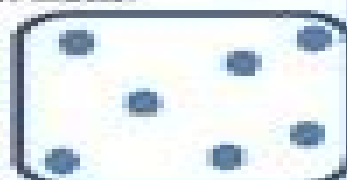
## Classification of Matter



## Phases of Matter

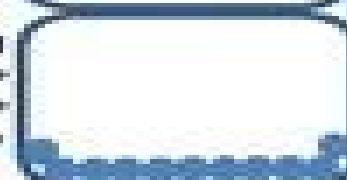
**Gas**

- NO definite volume
- NO definite shape
- Completely fills container
- Weak forces of attraction



**Liquid**

- definite volume
- NO definite shape
- flows and takes



**Solid**

- definite shape + volume
- regular geometric pattern
- vibrates in a fixed location



## Physical Separation of Mixtures

### Filtration

solid separated from liquid/solvents



key property:  
solubility (solid vs liquid/solvent)

### Distillation

liquids with different BP heated to boil

key property:  
Boiling Point

### Extraction

key property:  
intermolecular forces

### Chromatography

different color ink dye separated by solubility

key property:  
Molecular Polarity

## Significant Figures

### Atlantic Pacific Rule

Decimal point is present or significant



### Multiply/Divide

Answer must have least sig figs in data

$$\begin{array}{r} 24.68 \\ \times 3.250 \\ \hline 79.226 \\ \text{sig figs} \end{array}$$

# Chemistry Regents Survival Guide

**Effiong Eyo**



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